



MAGNETITE GRAPHENE FOR ELECTROCHEMICAL DETERMINATION OF URIC ACID

(Magnetit Grafin untuk Penentuan Elektrokimia Asid Urik)

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Received: 17 February 2019; Accepted: 22 May 2019

Abstract

In this work, the magnetite-reduced graphene oxide ($\text{Fe}_3\text{O}_4/\text{rGO}$) was fabricated for a selective voltammetric detection of uric acid. The magnetite-reduced graphene oxide is biocompatible sensing material that was synthesized through one-step reaction mechanism. The attachment of spherical like Fe_3O_4 particle on the rGO sheet was characterized by Fourier Transform Infrared (FTIR), Ultraviolet-Visible Spectroscopy (UV-Vis) and Scanning Electron Microscopy (SEM). The electrochemical performance of the $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ electrode were studied by Cyclic Voltammetry (CV) and Differential Pulse Voltammetry (DPV). The superparamagnetism property of the Fe_3O_4 anchoring on the rGO sheet provides a conduction pathway for the analyte diffusion meanwhile promote the electron transfer kinetic between the electroactive species and electrode surface. The enhanced current response signal produced from the reduction-oxidation process by the electroactive species were optimized further in parameters of scan rate, concentration of uric acid and pH value of the supporting electrolyte. The observed DPV response displays a linearity relationship on uric acid concentration ranging from $4\ \mu\text{M}$ to $100\ \mu\text{M}$ with correlation coefficient 0.989 and a sensitivity of $0.08478\ \mu\text{A}/\mu\text{M}$. The detection limit and limit of quantitation were found to be $3.04\ \mu\text{M}$ and $10.14\ \mu\text{M}$ respectively.

Keywords: magnetite, reduced graphene oxide, uric acid, electrochemical sensor

Abstrak

Dalam hasil kerja ini, magnetit-grafin oksida terturun ($\text{Fe}_3\text{O}_4/\text{rGO}$) telah difabrikasikan untuk tujuan pengesanan voltammetrik terpilih bagi asid urik. Magnetit-grafin oksida terturun merupakan bahan sensor bioserasi yang telah disintesis melalui mekanisma tindak balas satu langkah. Pelekatan zarah sfera Fe_3O_4 pada lembaran rGO telah dicirikan menggunakan Inframerah Transformasi Fourier (FTIR), Spektroskopi Sinar Ultra Lembayung Terlihat (UV-Vis), dan Mikroskop Imbasan Elektron (SEM). Prestasi elektrokimia elektrod $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ telah dikaji menggunakan kaedah Voltametri Berkitar (CV) dan Voltametri Pembezaan Denyutan (DPV). Ciri-ciri superparamagnetisma Fe_3O_4 yang terlekat di atas lembaran rGO menyediakan laluan konduktif untuk penyerapan analit dan juga meningkatkan kinetik pemindahan elektron di antara spesies elektroaktif dan permukaan elektrod. Peningkatan isyarat tindakbalas arus yang dihasilkan dari proses pengoksidaan dan pengurangan oleh spesies elektroaktif telah dioptimumkan lebih lanjut didalam pengoptimuman kadar imbasan, kepekatan asid urik dan nilai pH elektrolit sokongan. DPV yang telah diperhatikan menunjukkan hubungan linear di antara kepekatan asid urik antara $4\ \mu\text{M}$ sehingga $100\ \mu\text{M}$ berserta pekali kolerasi 0.989 dan sensitivity $0.08478\ \mu\text{A}/\mu\text{M}$. Had pengesanan dan had kuantifikasi masing-masing adalah $3.04\ \mu\text{M}$ dan $10.14\ \mu\text{M}$.

Kata kunci: magnetit, grafin oksida terturun, asid urik, sensor elektrokimia

Introduction

Uric acid (UA) is a biomolecule compound that plays significant role in regulate the homeostasis and metabolism system of the human body. The relative concentration of the UA varies dependently on tissue sites, which may accumulate temporarily or localized for extended period [1]. The UA that is sparingly water-soluble biochemical compound can be determined in the blood serum and excretion waste of the mammalian in the forms of urine. The clinical screening assessment for the uric acid serve as the fundamental tool in understanding various clinical conflicts that link to the uric acid level in human body. The abnormal amount of uric acid may potentially associated with diseases in term of chronic and acute that hold the risk of fatality [2]. The type of diseases that show close relationship to the abnormal amount of uric acid in body fluid including gout, hyperuricemia, rheumatoid and arthritis. Abnormal elevation in the level of uric acid in the body fluid will probably results several recognized clinical problem, which uric acid probably serve as the etiological agent. At high concentration, the uric acid can converted into oxidative radical and its derivatives, which has high cellular permeability and ability to interrupt during the DNA transcription and translation [3].

Quantitative and qualitative determination of the uric acid are significantly useful to provide related information especially in field of pharmaceutical, nursery, pathology, and medical field. There are varieties of techniques have been proposed with promising sensitivity and consistency in term of detection limit and the amount of uric acid quantified [4, 5]. Recently, electrochemical sensing has been highlighting as sophisticate alternative in identification of disease diagnosis molecules including UA [6]. Meanwhile, due to unique electronic properties and stiffness of the nano-sized graphene carbon allotrope, it has been widely adopted in material development and research activities was intensified on exploitation of graphene in photocatalysis, environmental remediation, and energy storage [1,7]. Graphene is a newly carbon material with nanoscale-dispersed phase at two dimensions other than zero-dimensional fullerene, one-dimensional carbon nanotube, and three-dimensional graphite [8]. The distinctive properties of graphene which are electrochemical beneficial have drawn noticeable attentions among the researchers to uncover its unique structure and properties. The graphene has 2D honeycomb lattice structure favored by the sp^2 hybridization of each carbon atom in the domain. The covalent reinforcement of the domain rendering graphene has strong mechanical strength and making it an ideal material for sustaining the engineering application as well as industrial manufacturing process.

Aside from its excellent physically strength, it has been found that graphene poses large surface-to-volume ratio, high thermal conductivity ($3000 \text{ W m}^{-1}\text{K}^{-1}$), and extremely high electron mobility at room temperature ($2.0 \times 10^5 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$) [8]. Such intriguing special of the graphene highly promoted as new potential electrode material for the application in electrochemically determination analyte of interest. A water-dispersible nanocomposite consisting of both magnetite and reduced graphene oxide was designed to address the issue associated with removal of arsenic in typical water treatment process [9]. By coupling both beneficial factors from the graphene and magnetite, the electronic and mechanical performance of the resulting compound may be enhanced. The combinations of excellent electrocatalytic activity of magnetite and good electronic properties of graphene serve as a promising hybrid material on measuring the UA content in biological sample. The UA monitoring examination is clinical diagnostic important as excessive amount of uric acid accumulated may potentially associated with disease of either chronic or acute as well as fatality.

The development of graphene-based hybrid material has been explore through its intended application in electronic industry, microprocessor chipset, and battery manufacture [10]. In order to resolve the difficulties coped in determination of uric acid in real sample, the fabrication of magnetite (Fe_3O_4) supported on the reduced graphene oxide (rGO) nanocomposite coated on glassy carbon electrode (GCE) is proposed in this paper. The $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ has demonstrated better biocompatibility as compare to bare glassy carbon electrode. Its noticeable electron transfer kinetic enhances the overall performance of the electrochemical instrumental about the desirable sensitivity and selectivity of electrode react with.

Materials and Methods

Reagent and apparatus

The concentrated sulfuric acid (H_2SO_4 , 98%), hydrochloric acid (HCl), ammonia (NH_3), potassium permanganate (KMnO_4), hydrogen peroxide (H_2O_2 , 30%) sodium nitrate and uric acid were purchased from the Merck, Germany.

The natural graphite powder, hydrazine hydrate, iron(II) chloride tetrahydrate ($\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$), and iron(III) chloride hexahydrate ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$) were obtained from Merck, Germany. The electrocatalytic performance of the magnetic nanocomposite decorated glassy electrode will be electrochemically analyzed by using 3-electrode system consists of counter electrode made of a platinum wire, silver/silver chloride electrode (Methrom AUTOLAB, Netherlands), with 3 M of KCl) as reference electrode, and magnetite-reduced graphene oxide modified glassy electrode (BASi MF-2012 3.0mm diameter) as the working electrode, with the use of Potentiostat/Galvanostat module PGSTAT (Metrohm AUTOLAB, Netherlands).

Synthesis of rGO

Firstly, GO was prepared by modified Hummer's method as reported by Yusoff et al. [11]. The graphene oxide solid sample obtained from the modified Hummers' method was grounded with mortar and pestle. The 0.5g of fine powder GO first dispersed in deionized water and sonicate for 30 minutes at 60 °C. Later, 5g of ascorbic acid added into the homogenized GO solution and stirred with magnetic bar for 30 minutes at 60 °C. The mixture underwent a color changes from yellowish brown to black precipitate, indicative of reduction of GO. The reduction product was treated with centrifugation and remove the supernatant. The black paste of rGO was added with excess 30% hydrogen peroxide to eliminate the unreacted AA by sonicating for 30 minutes at 60 °C. After sonication, the resulted black solid was obtained by centrifugation and wash with ethanol, deionized water 3 times respectively, and dried in the oven at 100 °C.

Co-precipitation of magnetite

The ferric (10.4 g) and ferrous chloride (4.0 g) can be prepared in a 2:1 mole ratio in 200 mL of degassed deionized water. The solution containing both ferric and ferrous species was heated up to 50 °C for 10 minutes. After the heating process completed, 2 mL of 25% ammonium hydroxide was added dropwise into the mixture solution under constant stirring rate at 50 °C. The pH of the mixture was adjusted using the relatively small amount of sodium hydroxide solution to maintain the solution within the pH range of 9-10. The solution stirred continuously at the same condition for one hour. The black solid magnetite was separated from the solution using a strong magnet, washed with 1:1 composition mixture of ethanol and deionized water. The solid mass was dried in the oven at 70 °C for 6 hours.

Preparation of magnetic nanocomposite ($\text{Fe}_3\text{O}_4/\text{rGO}$) through co-precipitation

The obtained graphite oxide powder (0.5 g) was homogenized in 300 mL deionized water through ultra-sonication for 35 minutes. The solutions of ferrous chloride (0.1584 g/25 mL) and ferric chloride (0.4055 g/2 5mL) will be prepared according to 1:2 molar ratio, which is the elemental stoichiometric need for the exact composition of the magnetite required. The resulting monodispersed solution was added with the as prepared iron (II/III) solutions slowly at room temperature and stirred at low speed power. Next, alkalization of the mixture solution by adding 30% ammonia solution. The pH of the mixture was adjusted to a magnitude of 10-11. Later on, the stirring mixture was gradually heated up to 90 °C and meanwhile promoting a reducing atmosphere by introduction of suitable amount of hydrazine hydrate (10 mL). The solution heated up for 4 hours at constant stirring condition. The resulting black solution cooled under room atmosphere and filtered with mixture of ethanol and water (1:1). The black powder obtained after drying in the oven at 70 °C for 2 hours.

Fabrication of the magnetite-graphene composite film

Prior to the surface modification of the bare glassy carbon electrode (3 mm in diameter), the surface was polished with fine alumina slurry (0.05 μm) and rinse thoroughly with deionized water. The GCE will then ultra-sonicate in the acetone and deionized water successively and dry under ambient condition. The $\text{Fe}_3\text{O}_4/\text{rGO}$ (1.0 mg) was added into 1 mL of deionized water and produced a homogeneous suspension after completion of one-hour sonication. Finally, 5 μL of the $\text{Fe}_3\text{O}_4/\text{rGO}$ homogenized solution was drop-casted on the pretreated GCE and allowed to dry in a vacuum desiccator for 2 hours.

Electrochemical Measurement

The $\text{Fe}_3\text{O}_4/\text{rGO}$ / GCE-modified electrode was used to determine uric acid by electrochemical means. The potential applied varies from -0.2 to 0.8 in order to study the redox behavior of the uric acid under cyclic voltammetry. The unmodified glassy carbon electrode was used to replace the $\text{Fe}_3\text{O}_4/\text{rGO}$ / GCE-modified electrode in comparison of

their thermal stability, electron kinetic transfer, electrode performance and efficiency. Other than this, the stability of the electrode designed will investigate at different external condition in the experiment. Parameters such as pH, concentrations and scan rate were used to study the optimum performance of the overall electrochemical system built.

Results and Discussion

Physical characterization of Fe₃O₄/rGO

The distinction in the functionalities and moieties of each constituents can be determined using the substrate nature and characteristic peaks demonstrated in the IR spectrum. The pure graphite in Figure 1(a) displays a smooth line passes through the region of 4000 to 1500 cm⁻¹ with no single peak appear. The layer intercalation and surface modification of graphite using the modified Hummers method was indicated successfully by the changes in the shape of the spectrum obtained for GO. The FTIR spectrum of GO shown a broad peak at the region around 2950-3699 cm⁻¹, which designated to the stretching vibration of OH groups of water molecules adsorbed on graphene oxide. The oxygen functionalities of the graphene oxide described by the peak value of 1730 cm⁻¹, 1400 cm⁻¹, 1210 cm⁻¹ and 1020 cm⁻¹ corresponding to the carbonyl C=O, carboxyl COO-, epoxy C-O, and alkoxy C-OH stretching vibration [12, 13]. The absorption band value of 1612 cm⁻¹ being assigned to the aromatic C=C stretching vibration of the carbon domain. Upon successful chemical reduction of GO with AA, the characteristic absorption peaks responsible for the oxygen moieties present in the GO no longer observed in the rGO spectrum [14–16]. A significant peak at 580 cm⁻¹ was determined in the spectra of synthesized Fe₃O₄ nanocomposite, where it defines the stretching vibrational mode of Fe-O intrinsically at the tetrahedral void [17]. The stretching vibrations of -OH at low-frequency region of 3400 cm⁻¹ and the band 1612 cm⁻¹ was ascribed to the bending mode of the water molecules absorbed on magnetite surfaces [18]. In the view of spectrum obtained for Fe₃O₄/rGO, the absorption bands being assigned to oxygenated group was greatly reduced in intensity and the transmittance band around 584 cm⁻¹ is attributed to the Fe-O bonding vibration, indicated the existence of magnetite [19]. There are additional symmetric and antisymmetric stretching mode of carboxylic group, COO- with corresponding absorption peak value at 1585 cm⁻¹ and 1414 cm⁻¹.

The optical properties of each constituent materials was measured and the progressive changes on the optical behavior were studied based on exposure upon the ultraviolet-visible light wavelength region. In Figure 1(b), the pure graphite exhibits no significant absorption peak in the range of 200 nm to 800 nm. This is mainly due to the non-functionalized surface and inertness nature of the pure graphite in response toward the radiation. The dramatic changes to the surface morphology and functioning group attached to the graphitic surface layer will promote the absorption of radiation that brought about excitation of outermost electrons into higher energy state. The GO displayed an optical π - π^* aromatic C-C bond absorption band at 231 nm and a shoulder peak at 300 nm which responsible to n - π^* transition of the carbonyl C=O bond [20]. Upon subjected to the AA green reduction, the as-prepared rGO revealed a bathochromic shift absorption at wavelength of 259 nm indicated the successful restoration of electronic conjugation structure and removal of oxygenated group covalently bonded to the graphene oxide surface [10, 21]. The Fe₃O₄ nanoparticle shows continuous absorption throughout the ultraviolet-visible range of 200 nm to 800 nm without any identifiable strong absorption peak. Similarly, the resultant Fe₃O₄/rGO demonstrated a sharp increment in the optical absorption particularly at near band edge followed by a steady state absorbance approximately at 268 nm [18]. The redshifted absorption pattern suggested the incorporation of Fe₃O₄ onto the surface of graphitic sp² domain of rGO which underlying an enhanced photocatalytic performance.

The surface morphology and structural feature of each type of constituents and nanohybrids materials were elucidated through SEM imaging. In Figure 2, GO and rGO demonstrated units of random aggregates with corrugation nature extending over the carbon domain. This unique surface changes was attributed to the chemical destruction and restoration of graphitic sp² hybridized network happened during intercalation stage and functionalization of graphitic surface [22]. The spontaneous graphitic sheet scrolling is thermodynamically favored, at which the irreversible wrinkle nature of rGO sustains a beneficial high ratio of total surface area to volume [23]. The existences of extra hole defects and vacancies spreading throughout the rGO sheets in comparison to GO is because of the introduction of sp³ oxygenated carbon population and discontinuation of the interplanar C-C bonding. This supported that GO was successfully reduced to rGO. The SEM image was obtained for magnetite nanoparticle as shown in Figure 2(d). The magnetite synthesized from co-precipitation appears as sphere-like

discrete particle with mean size differences in dimension of nanoparticle [24]. The resultant $\text{Fe}_3\text{O}_4/\text{rGO}$ nanocomposite was shown in Figure 2 (e) with visible discrete outline of both different rGO and MnP nanoparticles. The rGO film was densely anchored with spherical magnetite nanoparticle of uniform size. The crystal growth suppressive effect due interaction of GO sheet and Fe_3O_4 has restrained the extent of Fe_3O_4 agglomeration that could facilitate the electron transfer from Fe_3O_4 through the rGO sheet producing an enhanced electrochemical response.

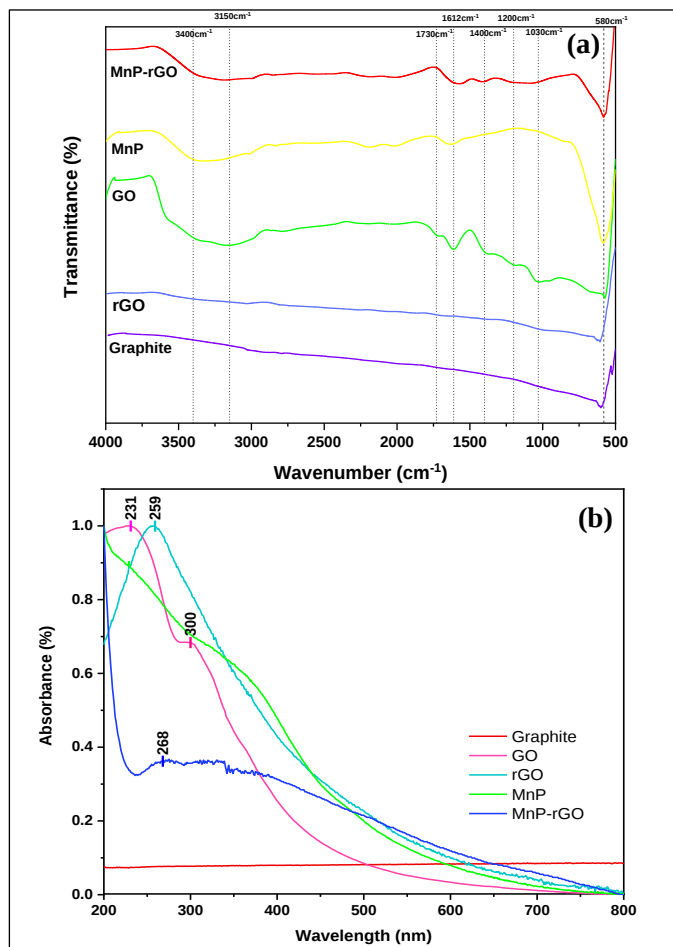


Figure 1. (a) FTIR spectrum and (b) UV-vis spectra of graphite, GO, rGO, Fe_3O_4 and $\text{Fe}_3\text{O}_4/\text{rGO}$ nanocomposite

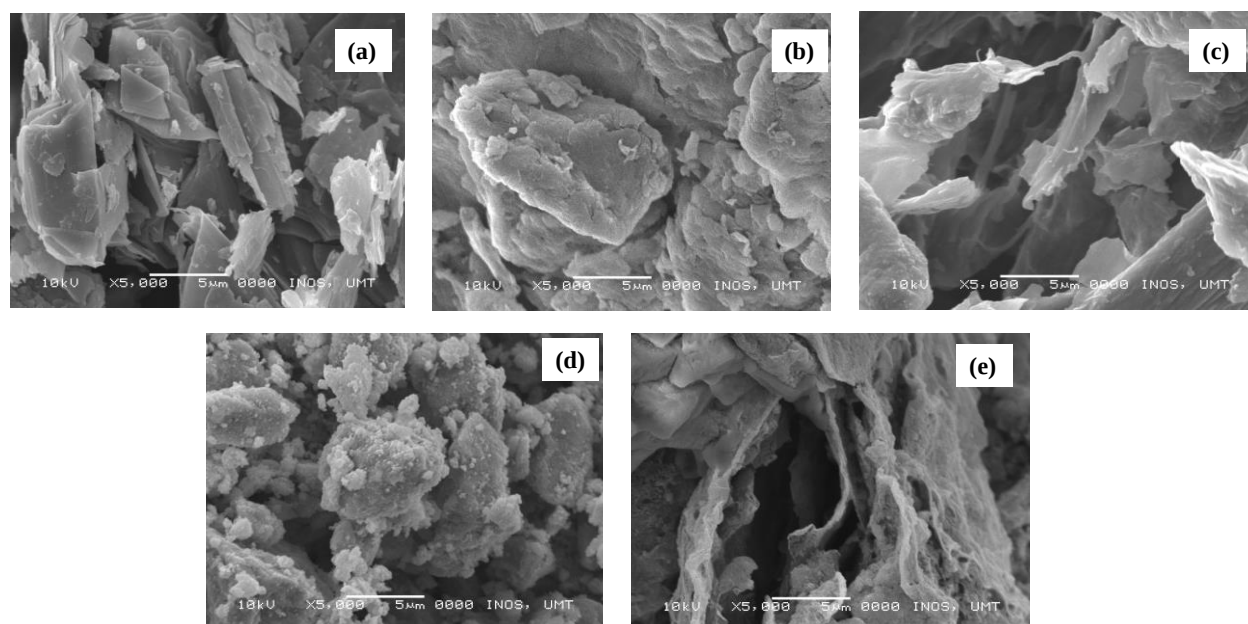


Figure 2. SEM micrographs of (a) graphite, (b) GO, (c) rGO, (d) Fe_3O_4 , and (e) $\text{Fe}_3\text{O}_4/\text{rGO}$ with magnification of x5000 and the operational voltage of 10 kV

Electrochemical characterization of $\text{Fe}_3\text{O}_4/\text{rGO}$

The electrochemical performance of $\text{Fe}_3\text{O}_4/\text{rGO}$ was studied using cyclic voltammetry as $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ modified electrode. For comparison purpose, bare GCE, rGO/GCE and $\text{Fe}_3\text{O}_4/\text{GCE}$ were also studied. Each individual contribution to the observed electrochemical responses were compared in the presence of supporting electrolyte, 1.0 M KCl and 5 mM of redox probe agent, $\text{K}_4[\text{Fe}(\text{CN})_6]$. The peak potential separation (ΔE_p) and the quasi-reversible behavior of the one-electron redox behavior of $[\text{Fe}(\text{CN})_4]^{3-/4-}$ used as indicative measurement in determination of electrode surface improvement. From Figure 3(a), CV measurement on the bare GCE electrode in the presence of $[\text{Fe}(\text{CN})_4]^{3-/4-}$ couple resulted both anodic I_{pa} and cathodic peak current I_{pc} at 391.4 mA and -384.4 mA, respectively. The ΔE_p measured for the bare GCE is 237.6 mV. After surface modification of GCE with rGO and Fe_3O_4 , the voltammetric response generated from both single Fe_3O_4 ($I_{pa} = 455.3$ mA, $I_{pc} = -433.3$ mA) and rGO ($I_{pa} = 524.2$ mA, $I_{pc} = -531.2$ mA) coating material which show remarkable peak current increment with a lower ΔE_p profile as 219.8 mV and 184.8 mV respectively. The rGO supported Fe_3O_4 hybrid composite further enhanced the electrochemical activity, where the anodic and cathodic peak current increase to 629.6 mA and -627.4 mA, almost two folds increase compare to bare GCE. This integrated hybrid material dedicated as promising platform of superior conductivity and electrocatalytic performance in sharpening the electrochemical measurement response for electroactive species which describes the $\text{Fe}_3\text{O}_4/\text{rGO}$ nanocomposites has the narrowest of ΔE_p (157.4 mV) among the other materials. The deviation from ideal ΔE_p value, 57 mV can be devoted to the incomplete chemical reversibility of redox couple, stability of analyte upon subjected to oxidation and subsequently reduced back to origin state, electron transfer kinetic rate, and existence of interfacial barrier [25]. On the basis of the slope measurement from Randles-Sevcik plot for ferrocyanide oxidation, the linear dependence of peak current, I_{pa} on the square root of scan rate can offer quantitative information about the effective surface area (EASA) and diffusion coefficient of the electroactive species. The proportional of Randles-Sevcik slope to the EASA was observed in the order of $\text{GCE} > \text{Fe}_3\text{O}_4/\text{GCE} > \text{rGO}/\text{GCE} > \text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ as shown in Figure 3 (b). The improved electroactive surface areas of $\text{Fe}_3\text{O}_4/\text{rGO}$ composite greatly highlights the excellent nature of surface layer deposited on GCE electrode overcome the performance achieved through rGO and Fe_3O_4 , individually. Data obtained from CV information is simplified in Table 1.

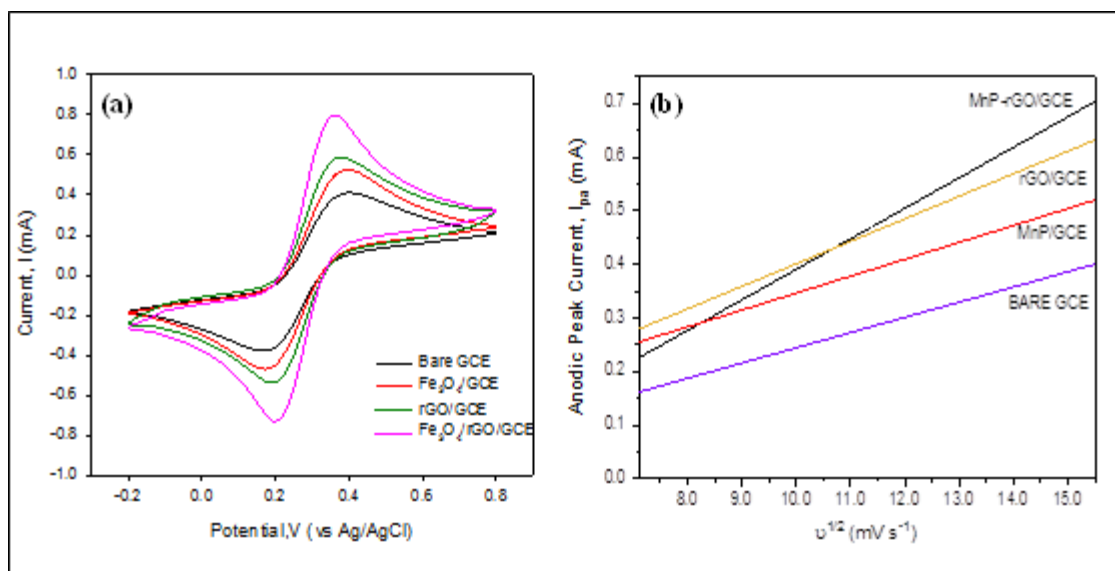


Figure 3. (a) Cyclic voltammogram of bare GCE, GO/GCE, rGO/GCE, $\text{Fe}_3\text{O}_4/\text{GCE}$, and $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$
(b) Randles-Sevcik Plot of different modified electrodes for EASA determination

Table 1. Peak potential separation for bare GCE and modified electrode in the presence of 1.0 M KCl and 5 mM redox probe solution

Electrodes	I_{pa} , μA	I_{pc} , μA	E_{pa} , mV	E_{pc} , mV	ΔE_p , mV	Slope (10^{-5})	EASA, cm^2 (10^{-6})
GCE	391.4	-384.4	409.9	172.3	237.6	2.84	8.48
$\text{Fe}_3\text{O}_4/\text{GCE}$	455.3	-433.3	394.8	175.0	219.8	3.16	9.42
rGO/GCE	524.2	-531.2	379.7	194.9	184.8	4.20	12.5
$\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$	629.6	-627.4	359.8	202.3	157.4	5.69	17.0

Electrochemical performance of $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ in the presence of UA

The voltammetric responses of the $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ toward UA detection and its application in different sensing environment was studied in the 0.1M PBS buffer solution of pH 7 containing 400 μM UA analyte. The electron transfer behavior through different sensing interface of electrode was studied using the same instrument configuration and parameter setups. The cyclic voltammograms of various modified electrodes in Figure 4 revealed the electrocatalytic performance of $\text{Fe}_3\text{O}_4/\text{rGO}$ nanocomposite coated GCE outstands the other surface-modified electrode. There was a minor difference in the anodic peak potential for the UA oxidation on various surface modified electrode. However, the rGO/GCE shows a higher current background baseline at the end of forward scan. This baseline signals was refined on the proposed $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$, which assumed due to the depleted electrical double layer (EDL), and capacitance current accumulated at the charged electrode. The anodic peak current has been improved as much as 1.65 μA (81.28%) compared to bare GCE under the same 50 mVs^{-1} scan rate. The increment in the anodic peak current, I_{pa} was parallel to the calculated diffusion coefficient using the Randles-Sevcik equation shown in equation 1.

$$I_p = (2.69 \times 10^5) n^{3/2} A D^{1/2} C v \quad (1)$$

where n is the number of electron taking part in the redox reaction ($n=2$), A is the electroactive surface area, cm^2 ($A = 0.08\text{cm}^2$ respect to GCE), D as the diffusion coefficient, $\text{cm}^2 \text{s}^{-1}$ ($c = 4 \times 10^{-4} \text{mol cm}^{-3}$) and scan rate, V s^{-1} ($v = 0.05 \text{V s}^{-1}$) is denoted by v .

The $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ modified electrode stands a better diffusivity ($D = 4.57 \times 10^{-13} \text{ cm}^2 \text{ s}^{-1}$) compared to another modified electrode. Thus, it do allows faster rate of mass transport for UA molecules diffuses across the concentration gradient and bulk solution molecules near the surface of electrode in a voltammetric detection. The cyclic voltammogram of UA in 0.1 M PBS buffer solution show a typical observation, where only one anodic peak was recognized. It expressing that the undergoing reaction is of irreversible for UA reduction. This can be attributed to the characteristic of UA and its possible alternative oxidation pathway of the electrode process demonstrated in the Figure 4 (inset) [26]. The primary partially oxidized UA product is composed of 4,5-dihydroxyluric acid, which is liable to oxidative degradation in favor the formation of chemically stable allantoin molecules. The diffusion coefficient results from bare GCE and different surface-modified electrodes summarized in Table 2.

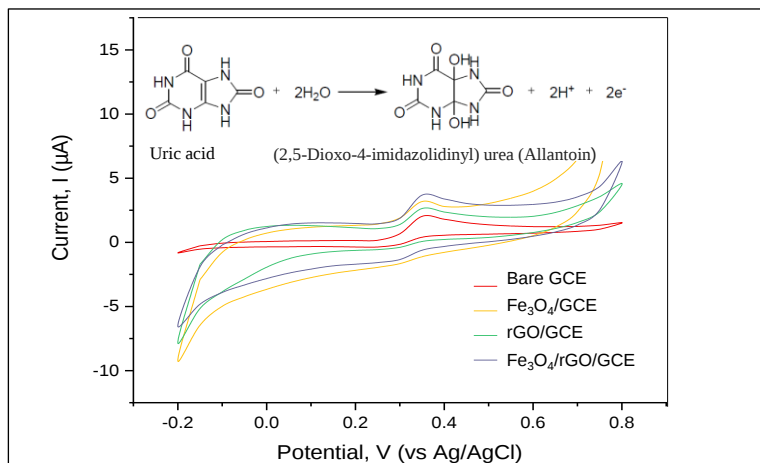


Figure 4. Cyclic voltammetric response of different surface-modified electrode in 400μM UA containing 0.1M PBS buffer solution (pH 7)

Table 2. The diffusion coefficient results from bare GCE and different surface-modified electrodes in 0.1M PBS solution (pH = 7) containing 400μM UA

Material (s)	Anodic Peak Potential, E_{pa} (V)	Anodic Peak current, I_{pa} (μA)	Diffusion Coefficient, D ($\text{cm}^2 \text{ s}^{-1}$)/(10 ⁻¹³)
Bare GCE	0.35	2.03	1.39
rGO	0.35	3.18	3.42
Fe_3O_4	0.35	2.64	2.35
$\text{Fe}_3\text{O}_4/\text{rGO}$	0.35	3.68	4.57

The optimum condition used to investigate the electrochemical behavior of UA was then performed based on the three parameters chosen for voltammetric detection purpose. The effect of scan rate, analyte concentration and pH of the supporting electrolyte were focused for any enhancement effect in the UA oxidation signal. The improved signal to noise ratio can be helpful in defining the UA content in the urine sample, which can be said equally important in recovery, precision and accuracy study that accounts for the repeatability as well as reproducibility. The conventional 3-electrode system was applied in parameter optimization with modified electrode $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ as the working electrode and a 2 mm diameter platinum rod dedicated as counter electrode in complement to silver/silver chloride, Ag/AgCl (KCl=3.0M).

The scan rate selected for cyclic voltammetry detection of UA ranging from 50 mVs^{-1} to 250 mVs^{-1} with a 50 mVs^{-1} as the increment constant at each running cycles ($n = 6$). The instrument calibration factor and diagnostic test was

performed before running the experiment. The electroanalysis profile selected in procedure and configuration was fixed with 5 mV step potential and starting potential of 0.15 V. The cyclic voltammogram displayed in Figure 5(a) revealed the effect of scan rate variation on the peak current recorded for the UA oxidation process. Table 3 depicted that increasing the scan rate from 50 mVs⁻¹ to 250 mVs⁻¹ positively shifts the peak potential from 0.356 V to 0.367 V. Correspondingly, the scan rate variation increases linearly with the diffusion coefficient, D calculated that facilitated the flux movement and diminishing of diffusion layer grow in an electrode process. This gradually shift in peak potential ascribed to UA oxidation implying the consequence of kinetic limitation happened between the active edges of the Fe₃O₄/rGO decorated on the GCE surface and UA molecules [24]. The linearity relationship between the peak current value recorded and variation of scan rate defined the characteristic of a surface adsorption-limiting process [27].

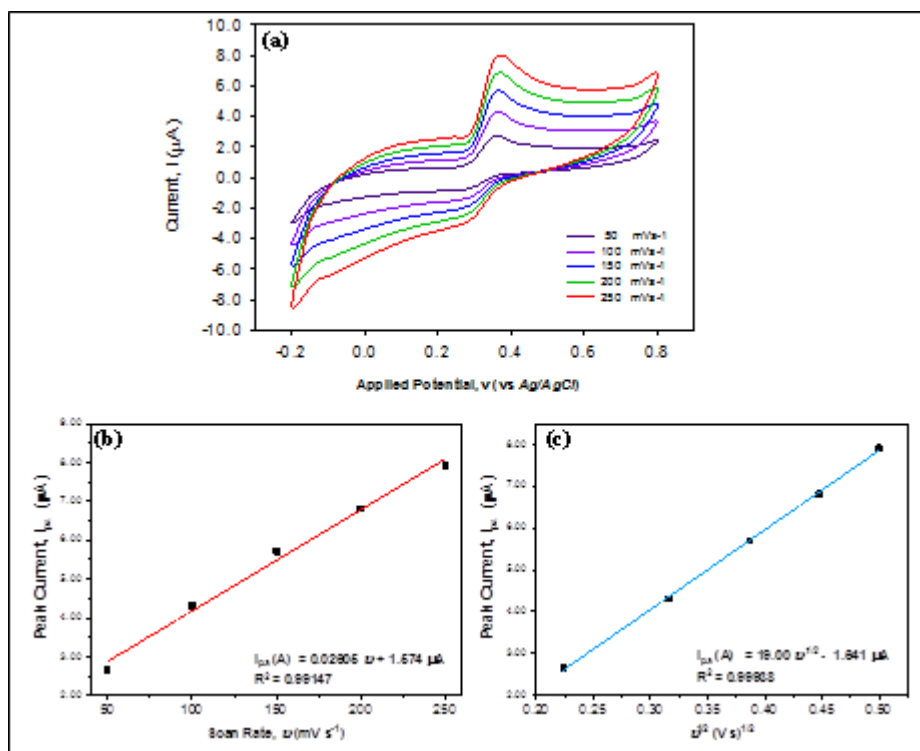


Figure 5. (a) Cyclic voltammogram of 400μM UA standard measured at different scan rate (50- 250mVs⁻¹) and (b) Plot of peak current against scan rate and (c) Randles-Sevcik plot for diffusion coefficient determination

Table 3. Scan rate variations and its related properties measurement result

Scan Rate (mV s ⁻¹)	Anodic Peak Potential, E _{pa} (v)	Anodic Peak Current, I _{pa} (A)/(10 ⁻⁶)	Square Root of Scan Rate (V ^{1/2} s ^{-1/2})	Diffusion Coefficient (cm ² s ⁻¹)/(10 ⁻¹³)
50	0.356	2.65	0.224	2.378
100	0.359	4.32	0.316	6.290
150	0.362	5.70	0.387	10.97
200	0.364	6.81	0.447	15.66
250	0.367	7.92	0.500	21.17

The Figure 5(b) displayed the peak current is directly proportional to the scan rate with a linear regression equation of $I_{pa} (A) = 2.605 \times 10^{-8} v + 1.574 \times 10^{-6}$, $R^2 = 0.99147$. Meanwhile, another identical relationship was illustrated in Figure 5 (c) which is a Randles-Revcik plot useful in effective surface area (EASA) and diffusion coefficient, D of a substrate molecule. A higher proportional linearity ($R^2 = 0.99938$) relationship was identified in the plots of anodic peak current (I_{pa}) against square root of scan rate variation ($v^{1/2}$) under the studied range of scan rate. The obtained regression equation $I_{pa} (A) = 1.9 \times 10^{-5} v^{1/2} - 1.641 \times 10^{-6}$ with correlation coefficient nearly to unity suggested the UA oxidation over the $Fe_3O_4/rGO/GCE$ surface was diffusion-limiting dominated over the adsorption-controlled mechanism [28]. These results provided that the mass transfer or absorption controlled process of UA oxidation take place at the low scan rate instead of diffusion-limiting dominated at high scan rate [29]. The scan rate 50 mVs^{-1} was selected for the further part of the optimization where its peak potential was nearly to the zero. In view of the current response of UA is being diffusion controlled, a slower scan rate will allow the UA molecules diffuses through the bulk electrolyte toward the electrode surface. Even though there will be growth in the diffusion layer, it provides certain degrees of sensitivity and stability about the variation in peak current and peak-to-peak separation especially in the study of pH parameter optimization.

A series of UA standard concentration was selected within range from $100 \mu\text{M}$ to $800 \mu\text{M}$. The experiment was designed to run 6 cycles simultaneously at 50 mVs^{-1} scan rate with a magnitude of 0.05 V step potential sweeping from initial potential -0.2 V to end potential 0.8 V . The reagent and standard preparations were kept constant as described previously. The step up current responses obtained from the increasing concentration of UA show a corresponding contribution of faradaic current from higher electroactive species in an electrode process. The respectively current response and peak potential variation according to the amount of UA introduced was tabulated in Table 4.

Table 4. Result of anodic peak potentials and peak currents at various UA concentrations (100-800 μM)

Concentration (s)	Anodic Peak Potential, $I_{pa} (V)$	Anodic Peak Current, $I_{pc} (\mu A)$
100	0.35	1.23
200	0.35	1.65
400	0.35	2.69
600	0.35	3.47
800	0.35	4.59

Based on the cyclic voltammogram result in Figure 6(a), the peak potential being unchanged for five successive different UA standard measurements at constant scan rate and experimental condition. This reflects good stability and accountability in measuring the current responses originated from various UA standard concentrations in suggested cyclic voltammetry technique. The calibration curve was plotted against the concentration of UA ranging from $100 \mu\text{M}$ to $800 \mu\text{M}$ provided a good fitting of data with correlation coefficient equal to 0.99711 . The linearity behavior of UA concentration and anodic peak currents can be expressed as $I_{pa} (A) = 4.757 \times 10^{-9} c + 7.270 \times 10^{-7}$, as shown in the Figure 6 (b). The highest current achieved was $4.59 \mu\text{A}$ correspondingly to the highest upper limit of UA concentration, which is $800 \mu\text{M}$ and the trend then followed in descending order of $600 \mu\text{M}$ ($3.47 \mu\text{A}$) < $400 \mu\text{M}$ ($2.69 \mu\text{A}$) < $200 \mu\text{M}$ ($1.65 \mu\text{A}$) < $100 \mu\text{M}$ ($1.23 \mu\text{A}$). The result obtained implies that the Fe_3O_4/rGO nanocomposite manage to electrocatalyze the UA oxidation meanwhile current transport at the electrode surface facilitated by gapless conducting material of rGO due excellent electron conductivity [1, 7]. The preferential concentration for the following research purpose was chosen $400 \mu\text{M}$, which is the intermediate value at full concentration range of UA studied. It gives sharp and well-defined oxidative current peak of considerable good sensitivity and peak resolution.

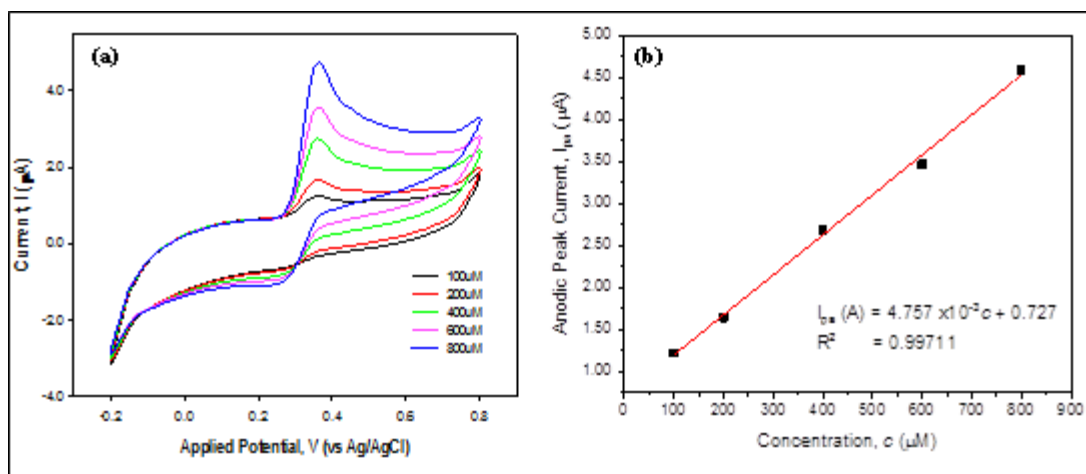


Figure 6. (a) Cyclic voltammetry measurements at various concentrations ranging from 100 μM to 500 μM of UA standard solution and (b) Calibration curve of anodic peak current against UA concentration (100-800 μM)

The effect of supporting electrolyte (PBS buffer solution) pH on the CV and DPV response of 400 μM UA were conducted by varying the pH from 4.0 to 9.0. The changes to the chemical environment surrounding the UA molecules in reaction of electrochemical oxidation was investigated in term of peak current and potential as well as the charging effect due electrolyte properties. Refers to Figure 7(a) and (b), the anodic peak potential shifts negatively as pH varying from 4.0 to 9.0 in both CV and DPV measurement. The DPV displays even higher sensitivity and selectivity as compared to typical CV technique due slow decay rate of capacitance current and non-faradaic current corresponds to the background current. These profound negative shifts suggested that the redox reaction of UA involved proton transferred at the Fe_3O_4 -rGO modified electrode. The pH changes will affects the adsorption performance of Fe_3O_4 /rGO, which can be rationalized based on the charge accumulation on the electroactive surface sites, degree of ionization of analyte, electrostatic repulsion or attraction between analyte and functionalized electrode surface [14, 30].

The uric acid solubilized in alkalizing solution presented majorly as deprotonated urate bearing an overall negative charged surface was repelled from redox site dominated by hydroxyl groups. Similarly, changes the pH from alkaline to acidic will neutralized the negative charged urate as non-ionized uric acid and the electrostatic attraction between the UA and Fe_3O_4 -rGO/GCE surface was strengthen [31]. Figure 7(c), the pH selected was 7.0, which is neutral and assembles to normal urine pH of 6.0 to 7.5. The peak current measured in both CV and DPV techniques revealed that the pH selected is stable and high sensitivity in view of high peak currents and negligible interference. The results obtained was summarized in Table 5. The plot of anodic peak potential versus pH in Figure 7(d) expressed linearity ($R^2 = 0.99509$) in the pH range of 4 to 9 with a slope of -56 mV pH^{-1} , which closely approximate to the theoretical Nernst equation value (-59 mV pH^{-1}). This again implies that the electrochemical oxidation of UA involves an equivalent number ($n=2$) of electron and proton [24, 31].

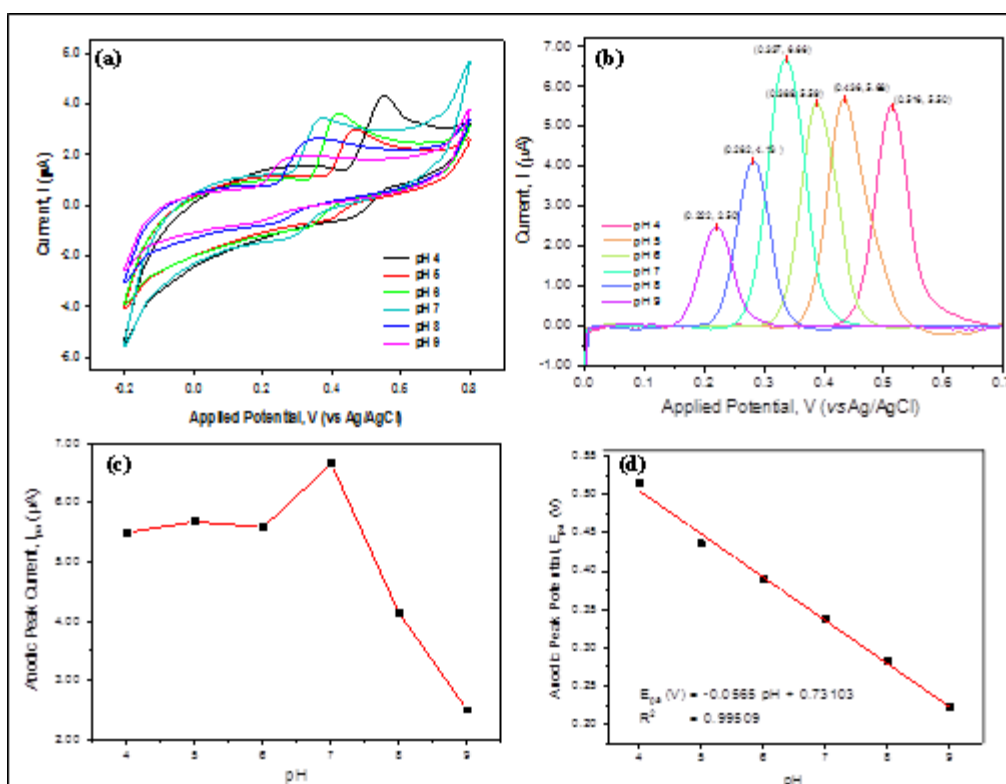


Figure 7. (a) Influences of electrolyte pH on the oxidative peak current of UA in cyclic voltammogram, (b) DPV response toward UA determination according to pH changes from 4.0 to 9.0, (c) Plot of peak current against pH value of the supporting electrolyte (0.1M PBS solution) and (d) Plot of anodic peak potential against pH value of the supporting electrolyte (0.1M PBS solution)

Table 5. Peak current and potential values correspond to the pH variation

pH	Cyclic voltammometric Response(s)		Differential Pulse Voltammetric Response(s)	
	Peak Potential, E_p (V)	Peak Current, I_p (μA)	Peak Potential, E_p (V)	Peak Current, I_p (μA)
4	0.553	4.34	0.515	5.50
5	0.470	3.03	0.436	5.68
6	0.424	3.64	0.389	5.59
7	0.373	3.48	0.337	6.66
8	0.337	2.57	0.282	4.13
9	0.262	1.79	0.222	2.50

Limit of detection (LOD) and Limit of Quantification (LOQ)

The LOD and LOQ calculated are $3.04 \mu\text{M}$ and $10.14 \mu\text{M}$, respectively, which is comparable, excellent and suitable for UA content determination compared to the previous study tabulated in Table 6.

Table 6. Comparison of current work with the previous study of uric acid determination

Electrode	Detection Method	Detection Limit (μM)	References
Pani/AuNPs	DPV	20.0	[32]
ZnO/rGO	DPV	8.5	[33]
AgNPs/rGO	DPV	8.2	[34]
Modified GCE (0.5M NaOH solution)	DPV	4.7	[35]
$\text{Fe}_3\text{O}_4/\text{rGO}$	DPV	3.04	Current work

Additionally, the current response proportional increases with the UA concentration added into the phosphate buffer solution of pH 7. In the lower range of concentration, the linear regression equation obtained was $I_{\text{pa}}(\text{A}) = 8.477 \times 10^{-8} c + 8.43743 \times 10^{-6}$ with a correlation coefficient of 0.989 and sensitivity of $0.08478 \mu\text{A}/\mu\text{M}$. In the second linear dynamic segment, the linearity was expressed as $I_{\text{pa}}(\text{A}) = 6.631 \times 10^{-9} c + 1.65937 \times 10^{-5}$ with a $R^2 = 0.953$ correspond to higher UA concentration ranging from $100 \mu\text{M}$ to $800 \mu\text{M}$ (Figure 8b). The declination in sensitivity (slope) at higher magnitude of concentration was described by the kinetic limitation of electroactive species toward the $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ surface in a transport mechanism [7, 36]. A surge amount of oxidation products inhibits the adsorption site for electrocatalytic oxidation of UA, thereby resulting lower down the electrode response sensitivity to $0.00663 \mu\text{A}/\mu\text{M}$.

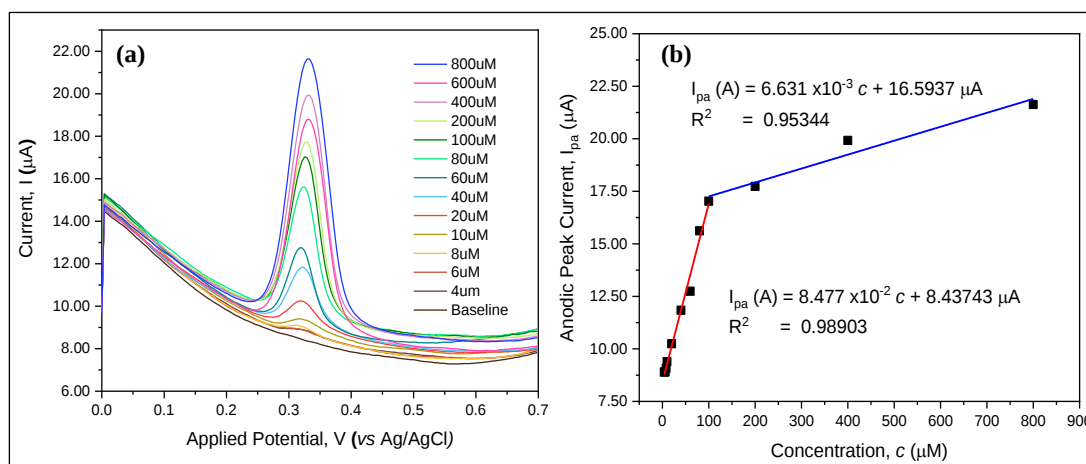


Figure 8. (a) DPV response measured over a range of UA concentration from $0 \mu\text{M}$ to $800 \mu\text{M}$ and (b) The relationship between the oxidation peaks current against UA concentration ($4\text{--}800 \mu\text{M}$)

Conclusion

The $\text{Fe}_3\text{O}_4/\text{rGO}$ hybrid material had successfully synthesized through one facile reduction of GO and co-precipitation of Fe_3O_4 which accounts for high percentage yield of product, cost effectiveness and being environmental green catalytic composite. The enhanced electrocatalytic performance of the $\text{Fe}_3\text{O}_4/\text{rGO}/\text{GCE}$ electrode provides a comprising potential in voltammetry determination of physiologically related molecules such as ascorbic acid and dopamine. The Cyclic Voltammetry (CV) result in the previous part has revealed the improved electronic properties of the resultant nanocomposite in term of the low resistivity, rapid electron transfer, and well defined peak resolution as well as peak separation. In additional, the physical properties of as-synthesized $\text{Fe}_3\text{O}_4/\text{rGO}$ nanocomposite was successfully investigated by various characterization techniques, including Fourier Transform Infrared Spectroscopy (FTIR), Ultraviolet-visible Spectroscopy (UV-Vis), Scanning Electron

Microscopy (SEM). The Differential Pulse Voltammetry (DPV) was used in complementary to the application CV and EIS in electrochemically characterization of Fe₃O₄/rGO nanocomposite. The Fe₃O₄-rGO/GCE sensing electrode exhibits high selectivity and sensitivity in UA oxidation with low detection limit of 3.04 µM. The results suggested the great potential use of Fe₃O₄/rGO/GCE in selectively non-enzymatic sensing of UA as compared to conventional UA determination such as chromogen-dependent colorimetric method and enzymatic technique by uricase.

Acknowledgement

The authors are grateful to Fundamental Research Grant Scheme (FRGS) FRGS/1/2017/STG01/UMT/02/2 for financial support as well as Universiti Malaysia Terengganu for providing facilities for undertaking this research.

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