



HYDROGEN SORPTION OF MAGNESIUM OXIDE CARBON NANOFIBRE COMPOSITE

(Serapan Hidrogen pada Komposit Magnesium Oksida Nano-gentian Karbon)

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Received: 13 April 2017; Accepted: 17 April 2018

Abstract

Carbon nanofibres have high specific surface area to adsorb hydrogen on their surface and are widely investigated for hydrogen storage. Although carbon nanofibres can store a considerable amount of hydrogen, the adsorption of the latter must be conducted at cryogenic conditions. Here, MgO is proposed as a catalyst to improve the hydrogen storage performance of carbon nanofibres at room temperature because of the light weight MgO and its ability to dissociate hydrogen molecules. The magnesium oxide-carbon nanofibre (MgO-CNF) composite was prepared with polyvinylpyrrolidone polymer and MgO via an electrospinner. The samples were characterised with field emission scanning electron microscopy, X-ray diffraction, Fourier transform infrared spectroscopy, thermogravimetric analysis, differential scanning calorimetry and microgravimetric analysis. The MgO particles were formed on the surface and embedded inside the MgO-CNFs, thereby increasing the specific surface area. The as-synthesised MgO-CNFs with a specific surface area of 547 m²/g can store 2.54 wt.% of hydrogen at room temperature, showing more than 30% improvement as compared with that of CNFs.

Keywords: electrospinning, hydrogen storage, carbon nanofibre, metal oxide

Abstrak

Nano-gentian karbon mempunyai kelebihan dengan luas permukaan yang tinggi untuk menyerap hidrogen di atas permukaan nano-gentian karbon, berpotensi sebagai bahan penyimpanan hidrogen. Walaupun nano-gentian karbon mempunyai kebolehan untuk menyimpan hidrogen yang banyak, penyerapan hidrogen perlu dilakukan pada keadaan kriogenik. Di sini, magnesium oksida telah dicadangkan sebagai pemangkin untuk meningkatkan prestasi nano-gentian karbon dalam penyimpanan hidrogen pada suhu bilik kerana ia adalah ringan dan berkebolehan untuk memisahkan molekul hidrogen. Komposit magnesium oksida nano-gentian karbon telah disediakan dengan polivinilpirolidon polimer dan magnesium oksida melalui kaedah elektroputaran. Sampel dicirikan dengan mikroskopi elektron pengimbasan pancaran medan, belauan sinar-X, spektroskopi inframerah transformasi Fourier, analisis termogravimetri, kalorimetri pengimbasan pembezaan dan mikrogravimetri. Zarah MgO terbentuk dipermukaan dan berada di dalam MgO-CNF telah meningkatkan spesifik luas permukaan. MgO-CNF dengan luas permukaan tentu 547 m²/g, berkebolehan menyimpan hidrogen sebanyak 2.54 wt.% pada suhu bilik, yang mana peratus peningkatan kapasiti penyimpanan hidrogen melebihi 30% berbanding dengan menggunakan CNF sahaja.

Kata kunci: elektroputaran, penyimpanan hidrogen, nano-gentian karbon, oksida logam

Introduction

In addition to their excellent thermal and chemical stabilities and fast adsorption kinetics [1], carbon nanomaterials, such as carbon nanofibres (CNFs) with a large specific surface area of up to $700 \text{ m}^2/\text{g}$, can adsorb hydrogen gas on the surface or in the pores [2]. Nonetheless, a reasonable hydrogen adsorption capacity of $> 2 \text{ wt.}\% \text{ H}_2$ [3] can only be achieved at cryogenic temperature or at high pressure. The recent study of Yang et al. [4] have shown that at 77 K and 1 bar, the hydrogen uptake of multiwalled carbon nanotube is only $0.21 \text{ wt.}\% \text{ H}_2$. Meanwhile, the maximum hydrogen storage of CNF obtained by Jaybhaye et al. [5] and Bai et al. [6] are 0.65 and $0.6 \text{ wt.}\% \text{ H}_2$ at 330 K and 9 MPa, respectively. Moreover, the amount of hydrogen absorption on all types of carbon materials at ambient temperature and pressure of up to 100 bars is less than $1.0 \text{ wt.}\% \text{ H}_2$ [7–9]. Hence, several attempts, such as the use of production techniques [10, 11], surface modification [12–14] and metal-doped materials [14–17] have been performed to enhance the hydrogen sorption of C-based materials. To exhibit a microstructure with a high surface area-to-volume ratio, CNFs can be produced using chemical vapour deposition (CVD), plasma-enhanced, arc discharge methods and pulsed laser deposition [10, 18, 19]. However, these methods are costly and difficult to produce in large scale [20]. One of the options to synthesise carbon nanofibre *via* electrospinning technique [21] because it is a versatile technique to produce micro- and nanoscale nanofibres. The CNFs produced from electrospinning have a small diameter in ranging from 5 nm to 500 nm [22], resulting in a large specific surface area.

Surface modifications *via* physical and chemical activations improve the hydrogen storage capacity [23]. Physical activation involves the gasification of carbon materials in the presence of oxidizing agents, such as CO_2 and steam. Meanwhile, chemical activation increases porosity *via* the chemical reactions between the reagents (KOH) and carbon materials. Chemical activation operates at a lower processing temperature and faster processing time, thereby producing higher porosity [23]. Activated carbon with a hydrogen storage capacity of $0.81 \text{ wt.}\% \text{ H}_2$ at 298 K and pressure of 100 bar was obtained by Kim et al. [24] after KOH activation. Nevertheless, a high specific surface area does not guarantee a high hydrogen storage capacity, especially at room temperature. Chen et al. [13] activated polyaniline-based carbon nanotubes with KOH and obtained a hydrogen storage capacity of $5.2 \text{ wt.}\% \text{ H}_2$ at a cryogenic temperature of 77 K and pressure of 50 bar. The chemical KOH activation of CNF was reported by Jiménez et al. [25], wherein the CNF adsorbed is $2.22 \text{ wt.}\% \text{ H}_2$ at 77 K and 10 bars. Nonetheless, storing hydrogen at cryogenic temperature is still less practical for an on-board hydrogen storage application. The presence of impurities, which reduces the hydrogen storage capacity [26], during the activation is also the main drawback of chemical activation.

Recently, researchers [27–30] have explored and investigated the use of metals on carbon nanomaterials to improve their ability to store more hydrogen at ambient temperature with low pressure. Metal and metal oxide such as NiO_2 , TiO_2 , CuO_2 and Pd have been commonly used as an additive in CNFs to explore their catalytic effect in improving the hydrogen adsorption capability of CNFs at room temperature. For instance, Im et al. [31] have achieved $2.8 \text{ wt.}\% \text{ H}_2$ at 100 bars and at room temperature by adding copper in activated CNFs. This result suggested that metallic compound can enhance the hydrogen adsorption via the spill-over phenomenon [3, 5, 8]. Furthermore, the surface modification with metal will create metal hydride formation. The hydrogen molecules adsorbed on the metal are dissociated, followed by migration of hydrogen atom migration on the carbon materials [14]. Some of the hydrogen atoms form stable bonds with surface, leading to irreversible adsorption. Lachawiec et al. [32] also stated that other hydrogen atoms can be reversibly released by migration back to the metal particles in the form of hydride. The favourable hydrogen uptake of metal-doped carbon nanofiber can achieve $2.81 \text{ wt.}\% \text{ H}_2$ at room temperature and at pressure of 70 bar [30].

In this work, magnesium-based material was chosen because of its light-weight characteristics, abundance and low cost as compared with NiO and CuO [33]. In addition, MgO has the catalytic ability to dissociate hydrogen molecules [34, 35], enabling it to possibly spill the hydrogen atom onto the carbon support. Hence, the combination of MgO and carbon nanofibre has a potential to improve the hydrogen storage capacity at room temperature with reduced pressure.

Materials and Methods

Materials

Both MgO (nano-powder size of < 50 nm) and polyvinylpyrrolidone (PVP) ($M_w = 1,300,000$ g/mol) were purchased from Sigma-Aldrich. Absolute ethanol (99%) was purchased from HmbG. All chemicals were used as received without further purification.

Preparation of the MgO–CNF composite

The precursor solution for electrospinning was prepared by mixing 7.50 g MgO and 2.50 g PVP in ethanol at 40 °C. The solution was stirred with a magnetic stirrer for 4 hours to obtain a homogeneous solution. Then, the homogeneous solution was electrospun through a 20 mL syringe with a capillary tip size with diameter of 0.45 mm, which was placed horizontally on an electrospinner (KSS Company), to produce the nanofibre composite mats. Electrospinning was performed at a voltage of 18 kV and at 15 cm between the tip and the rotating drum that was covered with an aluminium foil. A syringe pump was used to control the flow rate of the solution at 1.5 mL/hours. The nanofibre composite was stabilised at 80 °C for 2 hours before it was carbonised at 800 °C for 2 hours in nitrogen gas to form MgO–CNFs.

Characterization

The morphology and surface structure of the samples were studied with a Supra 55VP field emission scanning electron microscope (FE-SEM). The nanofibres were sputter coated with platinum to increase their electrical conductivity. Average diameters of nanofibres were estimated by averaging 60 diameter measurements in the FE-SEM images. Energy dispersive spectroscopy (EDS) was used to map and identify the element distribution in the nanofibres.

Bruker-D8 Advance X-ray Diffractometer with a Cu K α radiation source ($\lambda = 0.154$ nm) was used to identify the phase compositions and the crystallinity of the samples. The X-ray diffraction (XRD) patterns were recorded in the scan range of 2θ from 10° to 80°. To characterise the chemical bonding between MgO and PVP, we obtained the FTIR spectra by using NICOLET-380. The measurements were obtained in the range of 4000 cm^{-1} to 400 cm^{-1} .

A STA 449 F3 Netzch Thermogravimetric Analyser (TGA) equipped with differential scanning calorimetry (DSC) was used to study the thermal stability and the different rates of energy in the endothermic (uptake heat) and exothermic (release heat) conditions. The samples were heated from room temperature to 700 °C at a heating rate of 10 °C/min under a continuous nitrogen flow.

The pressure–composition isotherm (PCI) hydrogen sorption was studied with a Cahn Thermax 500 microgravimetric balance with sensitivity of 1 μg at 298 K. A pressure sensor (0–6.9 MPa) was used to measure the hydrogen pressure in the chamber. All samples were tested for their hydrogen storage from 0.1 MPa to 5.0 MPa at room temperature. Before the measurement, the samples were outgassed overnight until they reach a constant mass. The whole system was purged with Ar gas, which flowed through a molecular sieve and a cold trap to remove moisture and any impure substances. Hydrogen was charged step-wise, and the data were recorded when the mass reaches equilibrium. Hydrogen uptake was calculated after the buoyancy correction of the experiment.

Results and Discussion

The surface morphologies of both the PVP–CNF and MgO–CNF composites are illustrated in Figure 1. The surface of PVP–CNFs (Figure 1a) was smooth and bead-free as compared with that of the MgO–CNF composite (Figure 1b). The morphology of the as-received MgO with irregular sphere particles is shown in Figure 1c. The incorporation of MgO in the precursor solution had a significant influence on the surface morphology of the CNF composite. The MgO–CNF composite was rough with particles forming on the surface and entrapping inside the nanofibres. The diameter of PVP–CNFs was 481 nm, with a specific surface area of 413 m^2/g . Meanwhile, the diameter of the MgO–CNF composite was irregular with an average of 459 nm and specific surface area of 547 m^2/g . As shown in Figure 1c, the as-received MgO particles with an average particle size of 27.4 ± 2.6 nm were agglomerated because of the high surface activity, as suggested by Ferari et al. [36]. The average particle size of the as-received MgO particles agrees with the average particle size of the MgO nanofibre composite of 25.3 ± 4.8 nm.

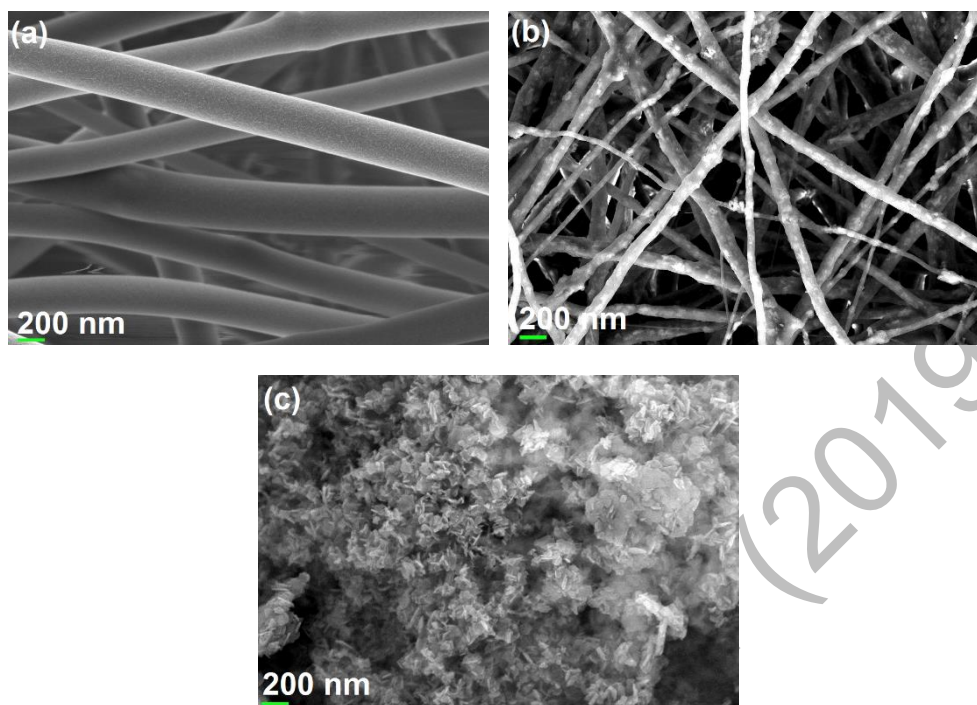
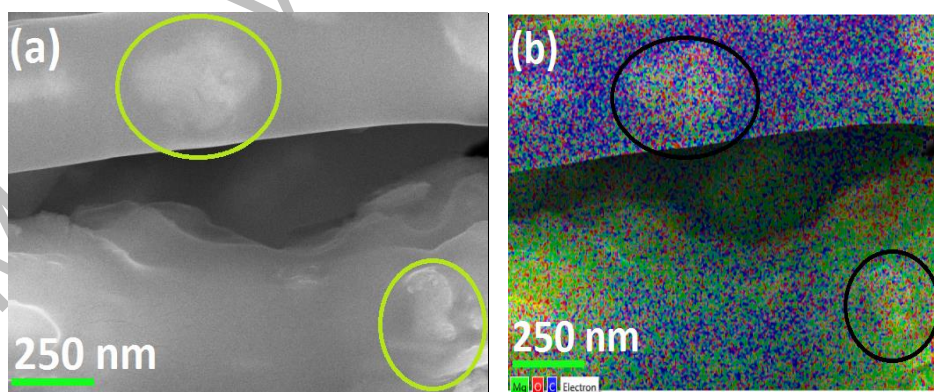


Figure 1. Field emission scanning electron microscope (FE-SEM) images of (a) the PVP-CNFs (b) MgO-CNF composite and (c) as-received MgO

EDS mapping was used to identify the elements in the CNF composite, as shown in Figure 2. The colour intensity of EDS mapping is associated with the concentration of elements in nanofibres; the higher the intensity, the higher the concentration and vice versa. The carbon element was uniformly distributed on the nanofibre, indicating that most of the other elements are embedded inside the CNFs. Meanwhile, high Mg concentration was detected at different areas of CNFs, corresponding to the particle-like bright spots (circle) in the SEM image. The results confirmed the presence of Mg compound in the CNFs.



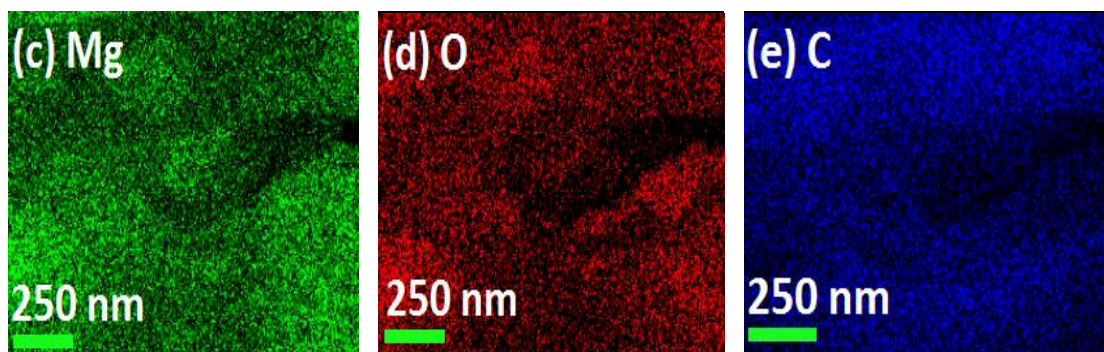


Figure 2. (a) Field emission scanning electron microscope (FE-SEM) image of the MgO–CNF composite. EDS mapping for (b) MgO–CNF composite, (c) Mg element, (d) O element and (e) C element

The XRD patterns of the PVP–CNFs, as-received MgO and MgO–CNF composite are presented in Figure 3. As shown in Figure 3a, two broad peaks are observed at 11° and 22° , which are induced by the PVP polymer. Both peak positions are similar with the results of Wang et al. [37] and Sivaiah et al. [38]. The broad peaks indicate that the CNFs have an amorphous phase. MgO peaks in Figure 3b were observed at the 2θ values of 37.1° , 43.1° , 62.5° , 74.6° and 78.6° , corresponding to the (111), (200), (202), (311) and (222) planes, respectively. All diffraction peaks were sharp, indicating that the MgO particles are crystalline [39, 40]. Additional peaks were observed at the 2θ values of 38.02° and 58.62° , belonging to the secondary $\text{Mg}(\text{OH})_2$ phase. These peak positions at the 2θ values of 38.02° and 58.62° correspond to the crystal (101) and (110) planes, respectively [41]. The formation of $\text{Mg}(\text{OH})_2$ was likely caused by the hydration of MgO upon exposure to humidity [42], where Mg^{2+} and OH^- were diffused into a layer of chemically and physically adsorbed water, forming and growing into a $\text{Mg}(\text{OH})_2$ crystal.

The XRD pattern of the MgO–CNF composite nanofibre is presented in Figure 3c. The pattern of the composite was a combination of Figures 3a and b, with no detected additional peaks, indicating that the Mg compounds are stable with CNFs. However, peaks belonging to MgO in Figure 3c slightly shifted to the lower angle as compared with the MgO phase in Figure 3b, suggesting that MgO may experience tensile residue stress caused by the pulling effect during the formation of nanofibres in the electrospinning process.

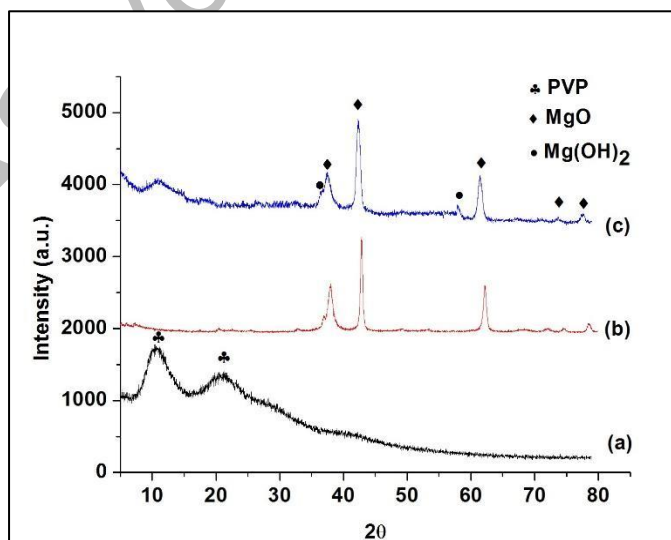


Figure 3. XRD patterns of (a) PVP–CNFs (b) as-received MgO and (c) MgO–CNF composite

FTIR was used to identify chemical bonding in the nanofibre composite. Figure 4 reveals the FTIR spectrum of the PVP nanofibre, PVP–CNFs, as-received MgO and MgO–CNF composite. Compared with the results of Baytak et al. [43], the PVP nanofibre spectrum had four main sharp peaks at 1659, 1463, 1421 and 1289 cm^{-1} in the present study. The peak at 1659 cm^{-1} is related to the stretching vibration of C=O. Meanwhile, the peak at 1463 and 1421 cm^{-1} are ascribed to the vibration bending and vibration deformation of methylene (CH_2) N-vinylpyrrolidone in the PVP chain, respectively. Another peak at 1289 cm^{-1} belongs to the stretching vibration of C–N. Other broad absorption peaks at 3420 and 2954 cm^{-1} are ascribed to the vibration bonding of O–H and the stretching vibration of C–H (methylene group), respectively [44]. All these peaks indicate the presence of PVP polymer in the nanofibre, which is consistent with the results of Dong et al. [45], Rafizah et al. [44] and Saravanan et al. [46]. After carbonization (Figure 4b), no significant characteristic peaks were observed, which confirms that only C atoms are present in the PVP–CNFs. These results are similar to the findings of Zhou et al. [47] and Guadagno et al. [48].

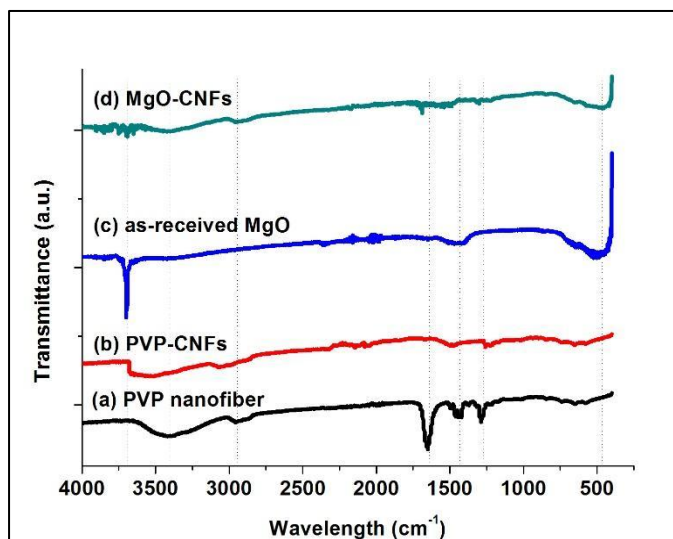


Figure 4. FTIR peak of (a) PVP nanofibre, (b) PVP–CNFs, (c) as-received MgO and (d) MgO–CNF composite

For as-received MgO, the MgO absorption peak was located at 490 cm^{-1} , which is similar with the results of Razaei et al. [25]. However, in the range of 4000–2500 cm^{-1} , a strong peak was detected at 3750 cm^{-1} . This peak was assigned to the isolated Mg–OH group. The presence of bending vibration of the O–H group is correlated to the exposure of the as-received MgO sample to humidity [51]. These results correspond to the observations in the XRD results. The FTIR spectra of MgO–CNFs were the superimposed FTIR spectra of both MgO and PVP–CNF. All functional groups and bonds corresponded well to their respective components, except that the transmittance signal of the Mg–OH group at 3750 cm^{-1} was less pronounced. This result indicates that the CNFs have changed the polarity of the Mg–OH bond.

The mass changes and thermal stability of the MgO–CNF composite are compared with those of the as-received MgO and PVP–CNFs, as shown in Figure 5. For TGA, a significant weight loss of the as-received MgO, PVP–CNFs and MgO–CNF composite can be observed at different temperatures as the temperature increases from room temperature to 700 °C. Meanwhile, the DSC graph shows multiple endothermic peaks (negative peak, releases heat). For PVP–CNFs, 5.2% weight loss was observed in the TGA curve between 25 °C to 166 °C, corresponding to a strong endothermic peak at 126 °C. The initial weight loss at temperature of less than 150 °C was due to the evaporation of solvent and any moisture from the nanofibre. This result is similar to the findings of Kim et al. [52] and Sales et al. [53]. At the temperature range from 373 °C to 466 °C, another significant weight loss of 71.5% was detected in the graph, which is associated with the thermal decomposition of the main chain PVP polymer [52]. Corresponding to the weight loss, an endothermic peak was also detected in the DSC graph at 449 °C. Beyond 500 °C, the PVP was decomposed completely, and a constant weight was obtained.

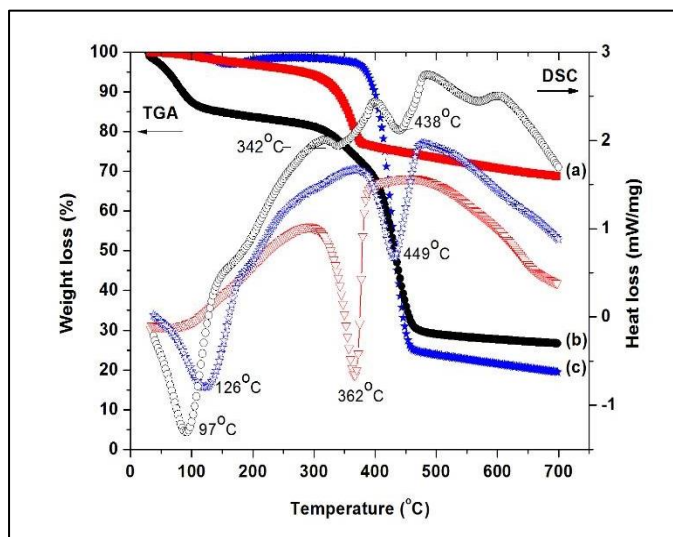


Figure 5. TGA–DSC graph of the (a) as-received MgO (▼) (b) MgO–CNF composite (□) and (c) PVP–CNFs (★). TGA: close symbol, DSC: open symbol

The as-received MgO showed a small weight loss of 0.98% between 52 °C to 311 °C. Similar to the results of PVP–CNFs, the first weight loss was related to evaporation of moisture that was trapped on the sample. A significant weight loss occurred at the temperature range from 326 °C to 374 °C, corresponding to an endothermic peak at 362 °C. This weight loss is associated to the decomposition of $\text{Mg}(\text{OH})_2$ [54], which is an impure compound that was detected in the XRD results. As the temperature increased above 400 °C, the weight of the as-received MgO was continuously reduced. This result is associated with the transformation from the amorphous to crystalline phase of MgO [55].

The presence of MgO in the CNFs had affected the thermal stability of PVP. Similar to the two other samples, the first weight loss of 14.03% accompanied a strong endothermic peak at 92 °C was caused by the loss of solvent and moisture from the nanofibre composite. The second weight loss of 56.22% occurred between 278 °C and 488 °C was similar to the decomposition of PVP–CNFs and MgO, corresponding to two DSC peaks, that is, two broad endothermic peaks at 342 °C and 438 °C. Both peaks shifted to a lower temperature as compared with its respective components. This result suggests that the addition of MgO lowers the decomposition activation energy of the MgO–CNF composite, as shown in the Kissinger analysis below.

Kissinger's equation [56] was used to estimate the decomposition activation energy of the MgO–CNF composite. The logarithm of heating rates ($\beta = 5$ °C/min, 10 °C/min and 15 °C/min) has a linear inverse relationship with the maximum desorption temperature at each heating rate according to the equation 1 below:

$$\ln [\beta/T_{max}^2] = -E_A/RT_{max} + A \quad (1)$$

where β is the heating rate, T_{max} is the maximum peak in the DSC curve, R is the gas constant and A is the linear constant. The activation energy was calculated from the slopes of $\ln [\beta/T_{max}^2]$ versus $1000/T_{max}$, as plotted in Figure 6. As shown in the figure, the incorporation of MgO into CNFs has lowered the decomposition activation energy of MgO from 191.53 kJ/mol to 182.68 kJ/mol.

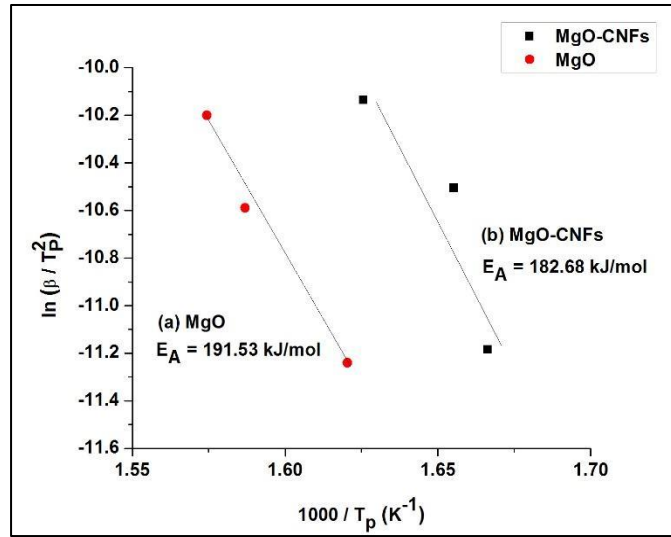


Figure 6. Kissinger's plot of the (a) as-received MgO and (b) MgO-CNFs

Figure 7 shows the hydrogen storage capacity of PVP-CNFs, as-received MgO and MgO-CNF composite at room temperature. PVP-CNFs, as-received MgO and MgO-CNFs adsorbed 1.90, 2.08 and 2.54 wt.% of hydrogen at 5.0 MPa and room temperature, respectively. With the addition of 75% of MgO in the CNFs, the hydrogen sorption capacity of CNFs was increased by approximately 30%, and the composite capacity was much higher than that of the as-received MgO. This improvement indicates that MgO has a synergetic effect in enhancing the sorption capacity because of a few possible reasons. Firstly, the rough surface morphology of the MgO-CNF composite increases the specific surface area, which provides more sites to adsorb the hydrogen. Secondly, MgO, which is distributed in the form of a single particle in CNFs homogenously in the composite, has effectively enhanced the hydrogen contact surface of MgO as compared with that of the as-received MgO, which tends to agglomerate. Lastly, the catalytic characteristics of metal oxide may reduce the adsorption activation energy and promote spill-over effect, as suggested by Yaakob et al. [30], who observed a similar improvement by using ZnO nanoparticles.

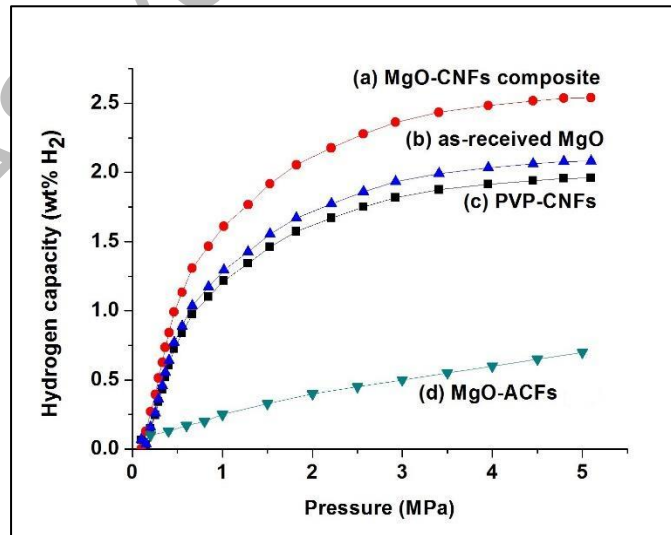


Figure 7. Pressure-composition isotherm of (a) MgO-CNF composite, (b) as-received MgO, (c) PVP-CNFs and (d) MgO-ACFs [31] at room temperature

The hydrogen sorption capacity of this work surpassed the findings of Im et al. [31], of which the MgO-activated polyacrylonitrile C fibres (MgO-ACFs) only adsorbed 0.7 wt.% H₂ at 5.0 MPa although the specific surface area of MgO-ACFs was up to twofold of that of this work. In this case, the role of MgO as a catalyst was suppressed probably because no MgO particles were exposed on the surface of ACFs, as indicated by the findings of Im et al. [3]. Presumably, the MgO particles on the surface have been removed during the chemical activation and washing process. Comparing the results, we can postulate that the position of MgO particles in the CNFs is of utmost importance to fully exploit the potential of metal oxide.

Conclusion

The MgO–CNF composite has been successfully synthesised using the electrospinning method. The addition of MgO in CNFs has increased the specific surface area of CNFs, which provides more hydrogen adsorption site and subsequently improves the hydrogen storage capacity. Nonetheless, the catalytic effect of MgO that promotes the spill-over mechanism has a larger role in enhancing the hydrogen storage capacity. Therefore, MgO must be located on the surface of CNFs to take advantage of its catalytic characteristics. This work has successfully improved the maximum hydrogen storage capacity of PVP–CNFs by more than 30% (2.54 wt%) at 5.0 MPa and at room temperature as compared with that of PVP–CNFs alone.

Acknowledgement

The authors would like to acknowledge the financial support given by the Universiti Kebangsaan Malaysia (DIP-2014-011, GGPM-2016-054) and Malaysian Ministry of Higher Education (FRGS/2/2014/ST05/UKM/03/1).

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