



THE ELECTROCHEMICAL BEHAVIOR OF ZINC OXIDE/REDUCED GRAPHENE OXIDE COMPOSITE ELECTRODE IN DOPAMINE

(Sifat Elektrokimia Elektrod Komposit Zink Oksida/Grafin Oksida Terturun dalam Dopamin)

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Received: 2 January 2018; Accepted: 6 March 2018

Abstract

Zinc oxide/reduced graphene oxide (ZnO/rGO) composite was synthesized via the modified Hummers' method. The presence of ZnO composite on the rGO sheets was confirmed by Scanning Electron Microscope (SEM), X-Ray Diffractometer (XRD) and Fourier Transform Infrared (FTIR) spectrometer. FTIR result showed the formation of ZnO/rGO with the presence ZnO stretching peak and the surface of ZnO/rGO was found to have better morphology with functionalized groups attached to it. XRD analysis also confirms the formation of ZnO/rGO composite by wurtzite structured indexed peak of the diffractogram. The ZnO/rGO composite was drop on the surface of GCE by drop casting method to increase the electrocatalytic activities of bare GCE. The electrochemical behaviour of ZnO/rGO modified electrode shows the enhancement in electron transfer and the system is diffusion-controlled. The electrochemical studies also revealed that the ZnO/rGO/GCE dramatically increased the current response against the dopamine (DA), due to the synergistic effect emerged between ZnO and RGO. This attributed that ZnO/rGO/GCE could exhibit excellent electrocatalytic activity and effective electron transfer kinetics towards the oxidation of DA. Therefore, the ZnO/rGO/GCE could be used to determine the DA concentration and provide an ideal matrix for clinical applications.

Keywords: zinc oxide, reduced graphene oxide, composite materials, dopamine, electrocatalysis

Abstrak

Komposit zink oksida/grafin oksida terturun (ZnO/rGO) telah disintesis melalui kaedah Hummer terubahsuai. Kehadiran komposit ZnO pada lapisan rGO telah diujikan oleh mikroskop imbasan elektron (SEM), Pembelauan sinar X (XRD) dan Inframerah transformasi Fourier (FTIR). Keputusan FTIR menunjukkan pembentukan ZnO/rGO dengan kehadiran puncak ZnO dan permukaan ZnO/rGO didapati mempunyai morfologi yang lebih baik dengan kumpulan berfungsi melekat padanya. Analisis XRD juga menunjukkan pembentukan komposit ZnO/rGO melalui difraktogram indeks puncak struktur wurtzit. Komposit ZnO/rGO telah diubahsuai dengan elektrod karbon berkaca (GCE) dengan menggunakan kaedah alas titis. Perlakuan elektrokimia ZnO/rGO/GCE menunjukkan peningkatan dalam pemindahan elektron dan sistem adalah proses serapan terkawal. Kajian elektrokimia mendedahkan bahawa respon arus ZnO/rGO/GCE meningkat terhadap DA, disebabkan kesan sinergi antara ZnO dan RGO. Ini disebabkan respon arus ZnO/rGO/GCE boleh mempamerkan aktiviti elektropemangkin yang berkesan terhadap kinetik pemindahan elektron DA. Oleh itu, ZnO/rGO/GCE boleh digunakan untuk menentukan kepekatan DA dan menjadi pemangkin sesuai untuk aplikasi klinikal.

Kata kunci: zink oksida, grafin oksida terturun, bahan komposit, dopamin, elektropemangkin

Introduction

Dopamine (DA) is the most significant catecholamines, belongs to the excitatory neurotransmitter family which consists of an amine group linked to a catechol structure. It has great influence on central nervous, endocrine, and cardiovascular system [1]. Moreover, DA plays a vital role in brain pleasure centers, emotional response and help to regulate movement for human. Abnormal level of dopamine in human body can lead to central nervous system disorders such as Parkinson's, Alzheimer's, Huntington's disease and schizophrenia diseases are cause by. Recent studies show that low dopaminergic activity, as indicated by high DA/metabolite ratios, was positively correlated with addiction [2]. Therefore, the development of an accurate and sensitive method for DA determination is necessary for the clinical diagnostic purposes.

Numerous analytical techniques are available for the detection of DA level such as UV-vis spectrometry, capillary electrophoresis, chemiluminescence and electrochemical analysis. Among these methods, electrochemical method are favorable owing to their advantages such fast speed, low cost, low detection limits and also high in accuracy [3]. However, uric acid (UA) and ascorbic acid (AA) are usually coexisted with DA in the extracellular fluids of the central nervous system in mammals or the biological fluid such as urine, blood and human breast milk [4]. Thus, AA and UA will become the interfering species that will result in poor selectivity and sensitivity towards DA detection.

In order to overcome this problem, two fundamental approaches were often performed; improve the electrocatalytic performance on the electrode surface to separate their oxidation potentials, and selective interaction with DA on the electrode surface without interferences [5]. From the strategies, various chemically modified electrodes have been fabricated using polymers [6], metal oxides [7], nanoparticles [8], and carbon based materials [9]. A great deal of attention has been given to the use of multiwall carbon nanotubes (MWNTs), and more recently to graphene, as modification materials on the surface of electrodes [10].

Graphene, a two-dimensional sheet of sp^2 -hybridized carbon, has receive a lot of attention in the electrochemical field due to its unexpected stability, exotic band structure and unique electronic properties, especially the high specific surface area and excellent electric conductivity low mass density. Graphene was obtained so far by several techniques including mechanical exfoliation of graphite [11], chemical vapor deposition [12], electrochemical reduction of exfoliated graphite oxide precursor [13] and chemical reduction of graphene oxide [14]. Among the various methods reported for reduction of GO, the chemical reduction of GO sheets pre-deposited on electrode surfaces seems to be a promising route for preparing reduced graphene oxide (rGO)-modified electrode surfaces because it is simple, high surface to volume ratio, abundant defect sites, and fast electron transfer rate [15]. One major drawback of rGO is the natural tendency of agglomeration and restacking due to van der Waals interactions, which lower the surface area of product [16]. To overcome this problem, modification of rGO with different functional molecules would have to be introduced to improve their electrochemical activity. Studies have shown that fabricating metal/metal oxide-nanocarbon hybrid have increased the catalysts properties by utilizing both the advantages of metal and nanocarbon [17,18].

In this research, the zinc oxide/reduced graphene oxide composite (ZnO/rGO) electrode has been developed to investigate its performance towards oxidation of DA. It was then used to determine the concentration of DA in human urine samples.

Materials and Methods

Chemicals and Materials

3-hydroxytyramine hydrochloride (DA), natural graphite powder, concentrated sulfuric acid (H_2SO_4 , 95 - 98%), hydrochloric acid (HCl), sodium nitrite ($NaNO_3$), potassium permanganate ($KMnO_4$), hydrogen peroxide (H_2O_2 , 30%), zinc nitrate ($Zn(NO_3)_2$), potassium hydroxide (KOH), potassium ferrocyanide ($[Fe(CN)_6]^{4-}$), potassium chloride (KCl), disodium hydrogen phosphate dodecahydrate ($Na_2HPO_4 \cdot 12H_2O$), and sodium dihydrogen phosphate dehydrate ($NaH_2PO_4 \cdot 2H_2O$) were purchased from Sigma Aldrich, USA. The stock solution of DA was prepared in the 0.1M of phosphate buffer solution. All reagents are of analytical grade and were used without any further purification.

Instrumentation

The electrochemical experiment was performed using Potentiostat/Galvanostat module PGSTAT30 (Metrohm AUTOLAB, Netherlands), three electrode systems consist of platinum wire counter electrode, glassy carbon electrode (BASi MF-2012 3.0 mm diameter), with or without modification of ZnO-decorated rGO as working electrode, and Ag/AgCl (Metrohm, with 3 M of KCl) as reference electrode. Physical characterization of ZnO/rGO were performed using Scanning electron microscope (SEM, JEOL JSM-6360LA), X-Ray Diffractometer (XRD, Rigaku MiniFlex II), and Fourier Transform Infrared Spectrometer (FTIR, PerkinElmer Spectrum 100).

Preparation of ZnO/rGO

The GO which synthesized from the graphite powder via the modified Hummers method [19] was reduced to reduced graphene oxide (rGO) by zinc nitrate as reduced precursors. Firstly, 0.01 mol of $\text{Zn}(\text{NO}_3)_2$ and 50 mL of distilled water were added to the conical flask and stirred vigorously with magnetic bar in the ice water bath for 1 hours. Next, 0.0904 g of GO was treated with 100 mL of distilled water to form GO solution. The GO solution was separated in to two centrifuge tubes (50 mL) with equal amount and ultrasonic for 1 hours. After 1 hours, the GO solution was added to the of $\text{Zn}(\text{NO}_3)_2$ mixture solution and stirred homogenously for 10 minutes. Then, 5.6 g of potassium hydroxide was added to the mixture solution and left for overnight. In this time, the pH for mixture solution is pH 8.5. After that, the mixture solution was centrifuged and filter by suction filtration. The distilled water was added during the filtration process to neutralize the rGO to pH 7. The rGO was dried at oven for 1 hour to obtain the ZnO/rGO precursors.

Fabrication of ZnO/rGO/GCE

The working electrode was prepared by dropping the ZnO/rGO solution onto a glassy carbon electrode. First, glassy carbon electrode (GCE) was polished by alumina oxide followed by rinsing with water, ethanol and water, in turn. To modify the ZnO/rGO/GCE, 1 mg ZnO/rGO powder was ultrasonic dispersed into 10 mL of distilled water for 20 minutes. Then, 8 μL of ZnO/rGO was dropped onto a clean GCE with a micropipette and dried at room temperature. The electrocatalytic ability of the ZnO/rGO modified electrode was investigated by cyclic voltammetry in 0.1 M of phosphate buffer solution (pH 7.0) as electrolyte.

Electrochemical characterization ZnO/rGO/GCE

The cyclic voltammetry of 5.0 mM $\text{K}_3[\text{Fe}(\text{CN})_6]$ in 0.1 M KCl on all electrodes were carried out at potential window of -0.2 to 0.8 V at room temperature with different scan rates using an electrochemical cell with three electrodes system. The electrolyte solution was purged with nitrogen gas to eliminate oxygen before use.

Results and Discussion

Characterization of ZnO/rGO: Fourier transform infrared (FTIR) analysis

FT-IR spectroscopy is known to be an important tool to characterize the graphene-based hybrid materials. In the FT-IR spectra of the pure graphite powder (Figure 1a), there is no significant peak observed in raw graphite. This attributed to that the graphite powder was pure and do not contain any functional groups on the graphite molecules. In the FT-IR spectra of GO samples (Figure 1b), there is a strong broad peak presence at 3390 cm^{-1} and several weak ones at 1728 , 1622 and 1053 cm^{-1} , which are attributed to the stretching vibration of O-H groups (carboxyl and hydroxyl), stretching of C=O, skeletal vibration of C=C (carboxyl or carbonyl), and stretching vibrations of C-O (epoxy or alkoxy), respectively.

The strong characteristics of C=O stretch at 1728 cm^{-1} indicating the present of oxygen-containing group in GO surface. After the GO is chemically reduced, the spectrum (Figure 1c) resulted in significant reduction of hydroxyl, carbonyl, and epoxide group. The C=O vibration band which located at 1728 cm^{-1} was disappeared, while the board O-H, C=C and C-O stretching bands remain. These results demonstrate that the functional oxygen-containing group on the GO surface were almost reduced after the reduction by $\text{Zn}(\text{NO}_3)_2$. Additionally, the spectrum for ZnO/rGO showed the distinct peak at around 507 cm^{-1} , which corresponds to ZnO [20]. Based on the analysis, one can conclude that GO was both reduced and functionalized by $\text{Zn}(\text{NO}_3)_2$.

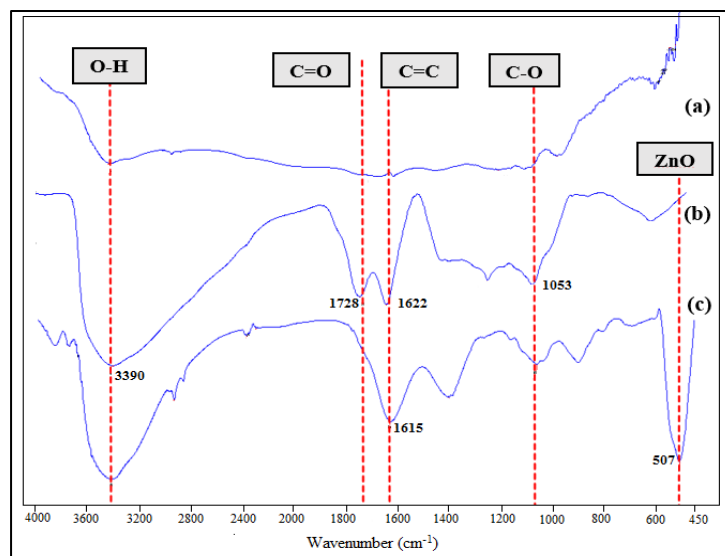


Figure 1. FT-IR Spectra of (a) Graphite, (b) GO, and (c) ZnO/rGO

Consequently, the purpose for reduction of GO was to eliminate the epoxy and hydroxyl groups on the plane, while other functional groups, such as carboxyl, carbonyl and ester groups which present at the edges of the carbon plane only have a minor influence on the conductivity of an rGO sheet. Moreover, the strong sharp peak of ZnO shown in FTIR spectra indicated that the ZnO were anchoring on the surface of rGO layer. Therefore, the electrochemical properties of ZnO/rGO composite is expected to enhance due to the wide band gap particles of ZnO and large surface area of graphene compared to those of pristine graphene oxide and ZnO particles alone.

X-ray diffractometry analysis (XRD)

The XRD analysis is used to verify the formation of the ZnO crystal on the surface of GO. The XRD pattern (Figure 2a) of graphite shows an intensive diffraction peak at $2\theta = 26.5^\circ$, reflecting the crystal form of graphite. If compare to GO spectra, the characteristic peak of GO at 10.25° was observed that indicates the presence of the larger interlayer distance between GO due to the oxidation of graphite and formation of the functional groups such as carboxyl, hydroxyl and epoxy, as indicated in the FT-IR results. While in the XRD pattern of ZnO/rGO, there are 8 peaks appearing at 2θ values of 31.7° , 34.3° , 36.1° , 47.3° , 56.4° , 62.6° , 67.7° , and 68.8° corresponding to the crystal planes of (100), (101), (102), (110), (103), (112), and (201) respectively. These peaks are indexed as wurtzite structured (crystal structure) of ZnO, which are consistent to the value in the standard card. Theoretically (002) peak of rGO is expected to appear in the XRD diffractogram of ZnO/rGO. However, it cannot be observed in the corresponding compositions (Figure 2b), indicating that the surfaces of rGO are fully covered by ZnO [21]. Hence, the diffraction peaks are mainly from the up-layer ZnO phase in the ZnO/rGO compositions. Moreover, the peak disappearance at 10.25° also indicating the surface of rGO is fully covered by ZnO. Therefore, it can be concluded that there is the formation of ZnO/rGO after the GO was reduced by zinc nitrate.

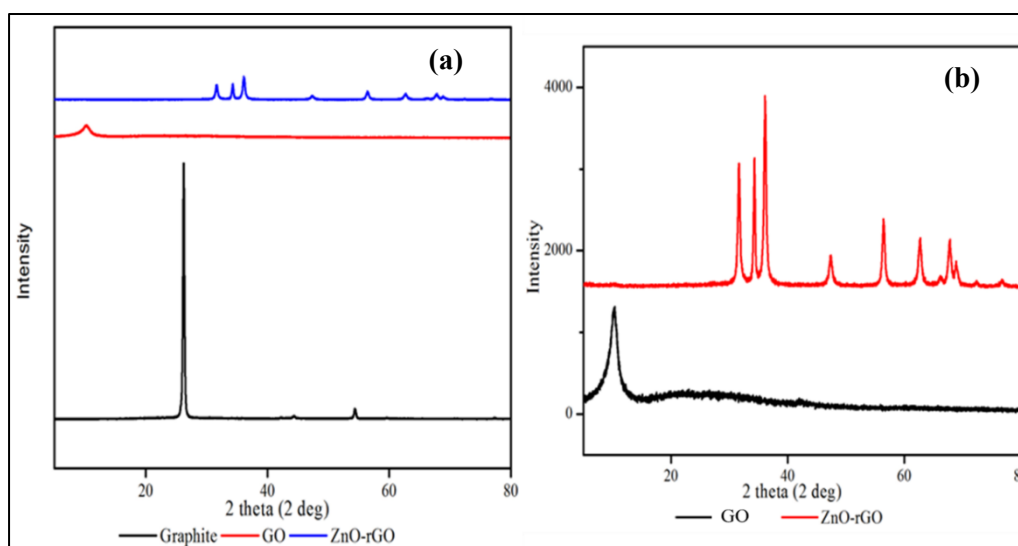


Figure 2. (a) XRD patterns for graphite, GO and ZnO/rGO composite and (b) Enlarge XRD patterns for GO and ZnO/rGO composite

Scanning electron microscopy analysis (SEM)

The morphology of the graphite, GO, and ZnO/rGO (Figure 3) was physically characterized by SEM. From the SEM image of pure graphite, the individual needle shaped particles are made up of graphite having intrinsic flake morphology. Graphite arranged in agglomerates and presented a typical multilayer structure. The SEM image of GO showing that the oxidation of graphite after treated with modified Hummers' method caused a visible change in their morphology, or it has a wrinkled appearance with multiple folds. This wrinkled aspect can be explained by the oxidation process to promote the admission of hydroxyl groups and epoxy groups, known as functional groups [21].

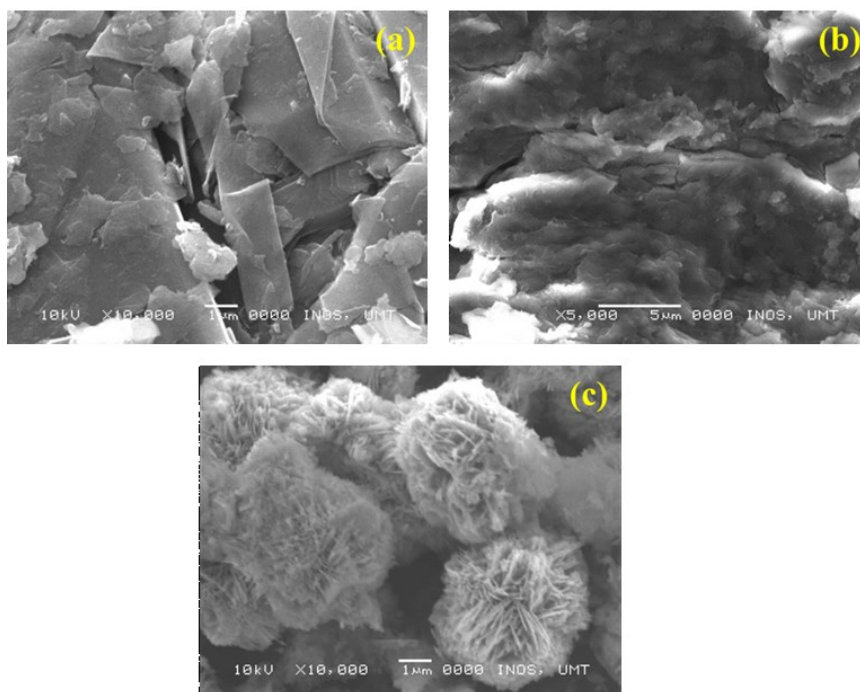


Figure 3. SEM image of (a) pure graphite; (b) GO and (c) ZnO/rGO composite

The formation of ZnO on the surface of graphene layer is indicated by the microsphere-like hierarchical structures after the GO is undergoes reduction by $\text{Zn}(\text{NO}_3)_2$. The ZnO composite with an average size of 5 μm were horizontally and randomly stacked between or on the RGO layers, in other words, graphene sheets are homogeneously incorporated in a ZnO matrix which clearly shown in Figure 3c. This result is also supported by XRD analysis. It may be due to the restriction of full length extension of ZnO in the space within the graphene layers. The results also indicated that the functional groups in GO were significantly reduced during $\text{Zn}(\text{NO}_3)_2$ reduction.

The electrochemical behavior of ZnO/rGO modified electrode

The electrochemical behavior of bare GCE, GO and ZnO/rGO/GCE was investigated by cyclic voltammetry (Figure 4a). Cyclic voltammetry (CV) is frequently used in analytical field due to it offers a wealth of experimental information and insights into both the kinetic and thermodynamic details of many chemical systems. The electrochemical behavior of bare GCE and ZnO/rGO/GCE is investigated by CV to elucidate their electron transfer behavior in the presence of 0.3 M KCl and 0.05 M $\text{K}_4[\text{Fe}(\text{CN})_6]$ at scan rate 50 mVs^{-1} and potential range in -0.2 V to 0.8 V. Curve of the bare GCE and ZnO/rGO/GCE showed a pair of well-defined quasi-reversible peaks with the anodic and cathodic peaks for bare GCE is at 348 mV and 200 mV, respectively. Meanwhile, the anodic and cathodic potential of ZnO/rGO/GCE are 298 mV and 190 mV, respectively.

The voltammogram shows the oxidation peak on ZnO/rGO/GCE was negatively shifted to 298 mV with remarkably enhanced of oxidation peak current. The ZnO/rGO/GCE present the largest background current and the peak current compared with the bare electrode. For the modified electrode (ZnO/rGO/GCE), the oxidation peak current was dramatically enhanced to 0.32 mA compare to GO and bare GCE. The decreased in peak potential separation ($\Delta E_p = E_{\text{Anodic Peak}} - E_{\text{Cathodic peak}}$) are observed for ZnO/rGO/GCE ($\Delta E_p = 99.8 \text{ mV}$) than that on GO and bare GCE electrode ($\Delta E_p = 150 \text{ mV}$).

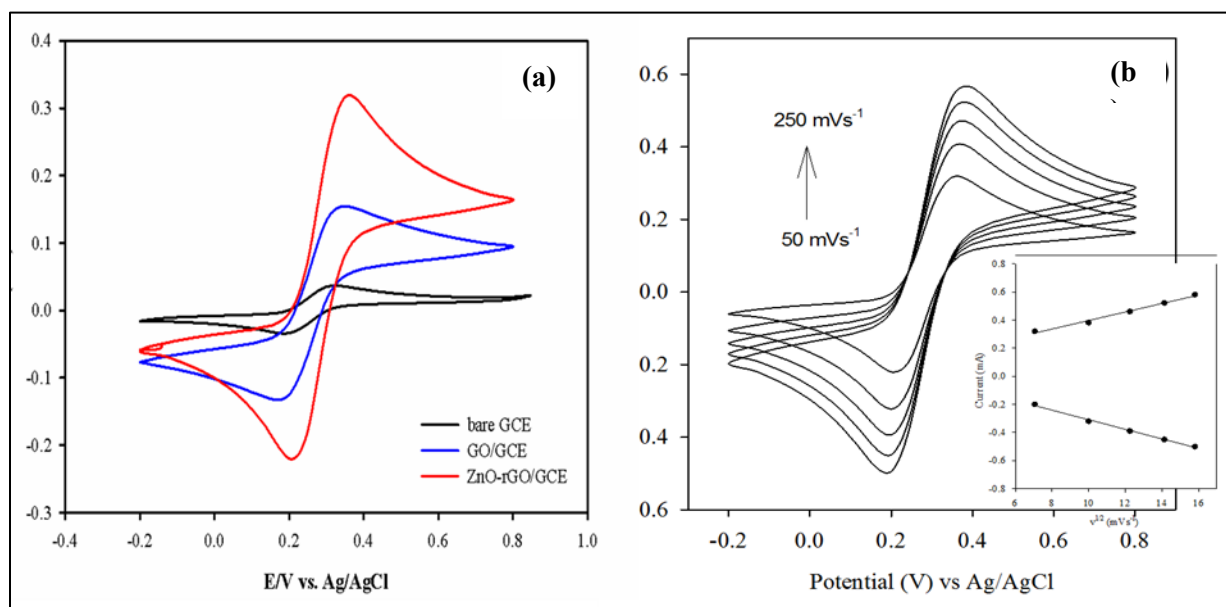


Figure 4. (a) Cyclic voltammograms obtained at the bare GCE, GO and ZnO/rGO/GCE in 0.3 M of KCl and 0.05 M of $\text{K}_4[\text{Fe}(\text{CN})_6]$ at a scan rate of 250 mVs^{-1} and (b) Cyclic voltammograms of ZnO/rGO/GCE at different scan rate. (Inset: The plot of square root of scan rate vs. current)

The response with smaller peak potential separation on ZnO/rGO/GCE was nearly reversible as a fully reversible system is $\Delta E_p = 59.2 \text{ mV}$ at all scan rate, which suggested a rapid electron transfer occurring at the interface of

ZnO/rGO/GCE. Besides, the redox peak currents increased obviously after the ZnO/rGO solution was drop-casted on the surface of bare GCE surface, which can be attributed to the synergistic effect of rGO and ZnO, promoting the electron transfer between the redox probe and the electrode surface.

The improved electrochemical behavior can be attributed to the excellent electrical conductivity of the ZnO/rGO present on the electrode surface. The carbon atoms in GO is decorated by oxygen-containing groups, which can expand the interlayer distance of graphene layer but also make the atomic-thick layers became hydrophilic. These properties will result the low electrochemical properties of synthesized GO [12]. Thus, the chemical reduction of GO should be carry out to reduce the oxygen-containing group which anchor on the surface of GO. Zn/RGO composite was produced after the GO was treated with $\text{Zn}(\text{NO}_3)_2$ as reduce precursors. The conductivity of monolayer graphene mainly relies on carrier transport within the carbon plane. Thus, various functional groups anchored to the plane are the main factor that influencing its conductivity.

The graphene with superior electrical conductivity and unique electron transport property could accelerate the electron transfer rate at the electrode/electrolyte interface. This shows that the electrical conductivity of ZnO is greatly enhanced the bare electrode. KCl and $\text{K}_4[\text{Fe}(\text{CN})_6]$ was used as the electrolyte to determine the electrochemical behavior for both electrode due to KCl is the inert supporting electrolyte which ensure that the ionic strength of the solution is high and hence that the electric field is homogeneous and near-zero and is not perturbed by the oxidation or reduction of the analyte concerned. While $\text{K}_4[\text{Fe}(\text{CN})_6]$ act as the redox species that present the reduction peak in cyclic voltammogram.

To investigate the reaction kinetics, the influence of scan rates on the electrocatalytic response of ZnO/rGO/GCE in $\text{K}_4[\text{Fe}(\text{CN})_6]$ was examined and the results are shown in Figure 4b. The scan cycles and potential range were standardized at value of 0.10 mM, 6 scan cycles, and -0.2 V to 0.8 V respectively. It can be clearly observed that the oxidation peak currents are increased linearly while the oxidation peak potential is shifted positively as the scan rate increasing from 50 mVs^{-1} to 250 mVs^{-1} , accompanied with an enlargement of the peak separation. Moreover, the increased of cathodic and anodic peak currents as the scan rate increased from 50 mVs^{-1} to 250 mVs^{-1} attributed that the electro-redox of ferrocyanide ion on ZnO/rGO/GCE is a typical diffusion controlled process [22]. In addition, at higher scan rates, the anode (E_{pa}) and cathode (E_{pc}) peak potential have a linear relationship with the square root of scan rate (Figure 4b: Inset). This is due to no stirring process was performed during CV test. Thus, the thickness of diffusion layer is increased as the time passes through and the concentration of ferrocyanide at the electrode surface had further diffused from the bulk solution. The faster the scan rate, the shorter the diffusion layer; hence greater the electron transfer [23]. Besides, the flux to the electrode is smaller at lower scan rate compare with the higher scan rate and the current is directly proportional to flux toward the electrode. Thus, scan rate 50 mVs^{-1} is selected as the optimize scan rate for further electrochemical study of DA at ZnO/rGO/GCE.

Electrocatalytic response of towards the oxidation of dopamine (DA)

Figure 5a shows the CV responses obtained for the bare GCE and ZnO/rGO/GCE in the presence of 0.1 mM DA in 0.1 M PBS by scan rate of 50 mVs^{-1} at pH 7. At a bare GCE, a well-defined redox peak is observed 347 mV and 243 mV with the anodic current of $3.39 \mu\text{A}$. However, at ZnO/rGO/GCE the oxidation peak current is dramatically increased to $20.0 \mu\text{A}$ which can be attributed to the enhanced electrochemical performance towards the oxidation of DA. This might be contributed from the excellent electrical conductivity of the ZnO/rGO presented on the electrode surface that can act as an effective electron promoter for electrocatalytic oxidation of DA. As the voltage increasing from -0.2 V, DA is attracted to the surface of ZnO/rGO/GCE. This situation lead the diffusion layer become greater as voltage increase, large amount of DA molecule adsorbed over the electrode surface.

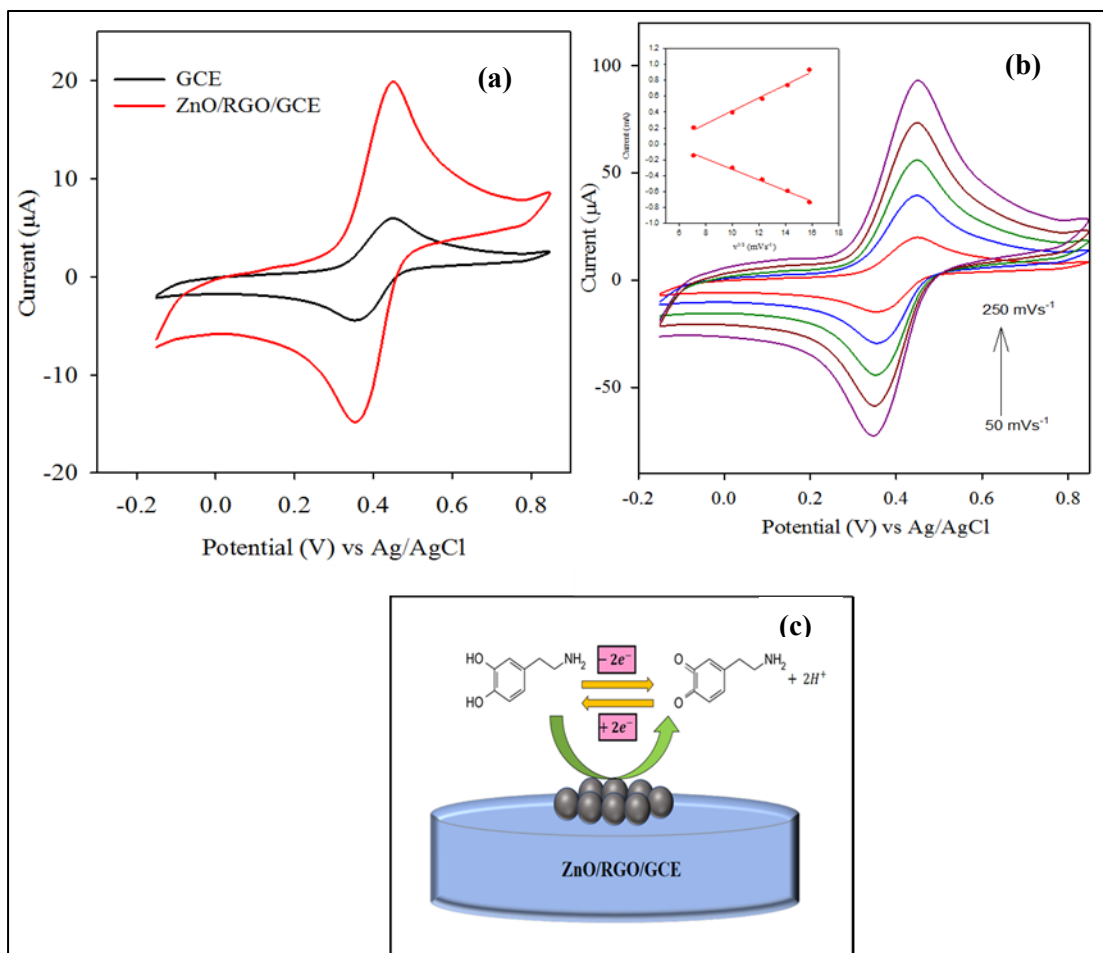


Figure 5. (a) Cyclic voltammograms obtained at the bare GCE and ZnO/rGO/GCE in 0.1 M of PBS (pH 7) containing 1.0 mM of DA, (b) Cyclic voltammograms obtained at ZnO/rGO/GCE in 0.1 M of PBS (pH 7) containing 1.0 mM of DA at different scan rate (50, 100, 150, 200, and 250 mVs^{-1}). (Inset: the plot of peak current vs. square root of scan rate) and (c) Electrocatalytic oxidation of dopamine at ZnO/rGO/GCE

The effect of potential scan rate was studied on the electrocatalytic properties of the ZnO/rGO/GCE toward oxidation of DA. As can be seen in Figure 5b, increase in scan rate potential, resulted in non-shifted potential value for the catalytic oxidation of DA. This result suggesting a stable kinetic in the reaction between the redox sites of the ZnO/rGO/GCE and DA. The result also shows that anodic and cathodic current increases with increasing of square root of scan rate indicating diffusion process. Redox reaction process of DA occurred at ZnO/rGO/GCE by two electrons transferred (Figure 5c). DA is oxidized during the positive sweep to form dopamine-o-quinone which is reduced back to dopamine in the negative sweep.

Effect of pH

The electrochemical behaviour of DA is strongly dependent on the pH value of electrolyte solution. Phosphate buffer solution (PBS) with different pH value (pH 3 to pH 11) at the surface of ZnO/rGO/GCE was prepared to investigate the optimum condition for the best electroactivity of DA standard solution. Figure 6 shows the characterizing of 0.10 mM of DA standard solution in 0.1 M of PBS with different pH at constant scan cycle (6 scan) and scan rate (50 mVs^{-1}) at ZnO/rGO/GCE. Cyclic voltammograms obtained with the ZnO/rGO/GCE for electrolyte containing DA in higher pH condition shows that DA did not couple with the ZnO/rGO/GCE and show

the lowest current peak compare to others. At pH 3 and 7 the oxidation peak current (I_{pa}) shows the maximum and obvious current. This indicated that the DA oxidized efficiency in both acidic and neutral condition. In this study, the redox reaction process of DA involves the transfer of two protons and two electrons. This phenomenon indicates that the uptake of electrons is accompanied by the same protons [24]. Therefore, considering the sensitive determination for DA, 0.10 M of PBS (pH 7) can be selected for further electrochemical investigation of DA at ZnO/rGO/GCE.

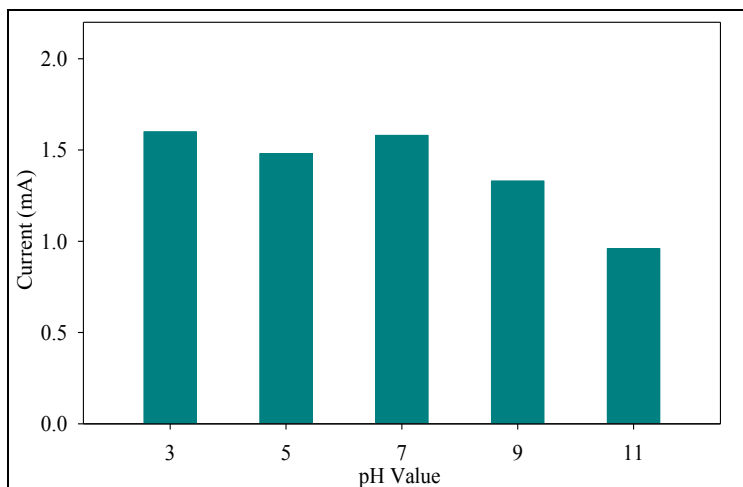


Figure 6. Chart of current vs. pH at the ZnO/rGO/GCE in 0.1 M of PBS with different pH value (pH 3, pH 5, pH 7, pH9 and pH 11) at scan rate 50 mVs^{-1} which containing 1.0 mM of DA

Conclusion

The modified electrode ZnO/rGO/GCE composite had been successfully synthesized via a simple, cost-effective and green approach method. The ZnO/rGO/GCE showed significantly improved peaks currents towards the redox reaction of DA compared with the bare GCE. Modified Hummers' method was proposed in this research to synthesized graphene oxide from raw graphite powder. Chemical reduction was carried out to reduce GO to ZnO/rGO. Chemical reduction was choosing due to the rGO which synthesize by this method is has a substandard surface area and electrically conductivity. The functional groups that anchor on the surface of graphene layer were characterized by Fourier transform infrared spectroscopy (FTIR). In addition, the morphology of ZnO/rGO was physical characterized by X-Ray diffractometer (XRD) and Scanning Electron Microscope (SEM). The FTIR results obtained indicate that the oxygen containing groups were almost reduced when the GO treat with ZnO/rGO as reduced precursor. Besides, the results from XRD and SEM proved that the ZnO composite were successfully covered on the graphene layer when the GO was chemical reduced. The ZnO composite which anchor on the surface of graphene layer will enhance the electrochemical behavior of GCE toward the redox reaction of DA. The ZnO/rGO/GCE displayed excellent electrocatalytic activity towards the DA oxidation. These results attribute that the ZnO/rGO/GCE could be a candidate for clinical, electrochemical and biosensor applications due to its excellent electrical performance.

Acknowledgements

The authors gratefully acknowledge to Ministry of Education for financial support from the Research Acculturation Grant Scheme (RAGS) RAGS/1/2015/ST0/UMT/03/2, Fundamental Research Grant Scheme (FRGS) FRGS/1/2017/STG01/UMT/02/2 and Central Lab UMT for providing facilities for undertaking this research.

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